This print-out should have 21 questions. Multiple-choice questions may continue on the next column or page – find all choices before answering.

001 10.0 points

An object undergoes a temperature change of $\Delta T=10^{\circ} C$.

What would be another way of expressing the same change?

- 1. A ΔT=10°C is equivalent to a ΔT=10°F.
- (2) A ΔT=10°C is equivalent to a ΔT=10K.

0002

- A ΔT=10°C is equivalent to a ΔT=50°F.
- A ΔT=10°C is equivalent to a ΔT=283.15K.

Convert 610°C	to Fahrenheit.
1, 2030 2, 2660	Tc = 1.8Tc + 32
3. 3650	= 1.4(010) +32
4. 2300 5. 1130	TF = 1130 F

10.0 points

5. 1130 6. 680 7. 1670 8. -58 9. 3200 10. 3830 Answer in units of °F.

9. 12.2564 10. 17.05

003 10.0 points

A building made with a steel structure is 610 m high on a winter day when the temperature is 0°F.

How much taller is the building when the temperature is $106^{\circ}F$? The linear expansion coefficient of steel is $1.1 \times 10^{-5} (^{\circ}C)^{-1}$.

80
*c
5

= 139.5144 m

AL= 2 L STC - CRUMINIM

9 1

Pressure

. Remains the same

Answer in units of cm.

004 10.0 points

Under constant pressure and with a constant amount of gas present, what happens to the volume of the gas if the temperature increase. Constant

- 1. Increases at first, then decreases
- 2. Increases PV = constant
- 3. Decreases
- 4. Remains the same
- 5. Decreases at first, then increases

005 __10.0 points

What is the volume of 2.6 mol of an ideal gas at a pressure of 3 atm and a temperature of 0° C? 1 liter = 0.001 m³ and 1 atm = 101300 Pascals.

scals. 1. 14.9373 2. 28.3809	Y= NKT (2.4)(6.3)	()(173k)
3. 22.4059 4. 24.6465	P 3 63 70	000 L
5. 15.6842	V = 19.4185 L	

- 6. 35.8495 7. 27.634
- 8. 33.6089
 9. 17.9248
 P=3am 101300 Fa = 30390 Fa 1L
 10. 19.4185
 Answer in units of (L)

006 10.0 points

The absolute temperature of a sample of monatomic ideal gas is doubled at constant volume. **PV= NRTP**

What effect, if any, does this have on the pressure and density of the sample of gas?

$$P = \frac{m}{V}$$

Density

Remains the same

- 2. Doubles Remains the same
- 3. Remains the same Doubles

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David J. Chalmers

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