

Gas Laws Lab Sheet

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Pre lab

1. After reading through the lab processes and pre lab section, explain the relationship between:
 - a. pressure and volume
The volume of a fixed amount of gas is going to be inversely proportional to its pressure at a constant temperature.
 - b. pressure and number of moles
If the temperature and the pressure are both constant, then the volume and moles are proportionate.
 - c. pressure and temperature
If the volume of gas remains constant, the temperature and pressure will directly change. When pressure is increased, temperature is also increased.
2. What causes pressure inside a container based on your review of the lab processes and pre lab materials?
The gas particles collide with the walls of the container which causes the pressure.
3. Complete the table below converting the following Celsius temperatures to Kelvin and Kelvin to Celsius. Please show your work.

Kelvin Temperature	Celsius Temperature
0	$0 - 273.15 = -273.15$
$0 + 273.15 = 273.15$	0
50	$50 - 273.15 = -223.15$
$75 + 273.15 = 348.15$	75
100	$100 - 273.15 = -173.15$

Data Collection

1. What happened to the pressure as the plunger was moved to create more or less volume?
When the plunger is moved to create more volume, the pressure decreases. When the plunger is moved to create less volume, the pressure increases.

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Ensheng Dong



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