

2.1.4 & 2.1.5 - Acids and redox

What are standard solutions?	A solution for which its concentration is known accurately.
What 3 things should you ensure when making standard solutions?	<ol style="list-style-type: none"> 1. Ensure you have clean dry equipment. 2. Make washings to remove any salt BEFORE MAKING up the solution to the specified amount. 3. Invert the volumetric flask to ensure everything is dissolved.
How does ammonia react with acid?	<ul style="list-style-type: none"> • ammonia + acid \rightarrow ammonium salt • $NH_3(aq) + HNO_3(aq) \rightarrow NH_4NO_3(aq)$
What are strong and weak acids ? Give 3 strong acids and 1 weak acid	<ul style="list-style-type: none"> • Strong acids dissociate fully to release H^+ ions in a solution e.g. $HCl \rightarrow H^+ + Cl^-$ • Weak acids dissociate partially to release H^+ ions in a solution – exist in equilibrium • Strong acids include: HCl, H_2SO_4 and HNO_3. • Weak acids include carboxylic acids such as CH_3COOH.
What are strong and weak bases? Give 1 weak base	<ul style="list-style-type: none"> • Strong bases dissociate fully to release OH^- ions in a solution • Weak bases dissociate partially to release OH^- ions in a solution – exist in equilibrium e.g. NH_3
What are diprotic and triprotic acids (with examples) and what does this mean for neutralisation?	<ul style="list-style-type: none"> • Diprotic acids and triprotic acids donate two and three protons respectively. • Examples include: H_2SO_4 and H_3PO_4. • They need to react to the number of moles of base to be neutralised

Unofficial Mark Scheme Chemistry 2013 Ocr

Michael Seilmaier



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