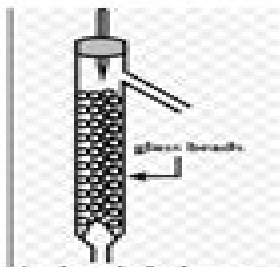


Suggest why chloroethanoic acid is stronger than ethanoic acid	(this question is already discussed a little further down.)		
fill in rate table	1.00 0.16 1.35 I beg someone write what the values were if they remember	K=250 Yes i agree - needs to be 1.4 as all values were to two sig figs Weren't we they to 3 dp though? The rate of the first one was 1.00 so I'm guessing 3sf of 2dp	3
empirical formula	C6H10O3S2	Percentages dudbt add up to 100 fuck me Didn't even see the sulphur, fml same same i thought i meant i was in the fuel area. YAY i got this!!	3
the labeling distillation	for the tube it was fractionating column or glass beads the bottom of the condenser was water in and top was water out	 <p>reject anti bumping granules might accept silica BEADS</p>	3
Draw a line for the reaction with chloroethanoic acid Suggest why chloroethanoic acid is a stronger acid than ethanoic acid.	<p>Was the line of chloroethanoic acid to the left of the first line, but levelled off at the same point? Yes but it levelled out before, same height though. Ah ok.</p> <p>Dissociates more Chlorine is more electronegative than hydrogen so will withdraw electron density from the carboxyl group meaning that the O-H bond breaks more easily dissociating into H⁺ making the solution more acidic</p>	Cl is electron withdrawing. it polarises the bond, weakens the O-H bond which releases more H ⁺ than ethanoic acid Neg inductive effect? Chloroethanoic acid forms a more stable conjugate base?	3
NMR Question?	<p>Use the integration values and the ratios to identify the structure at two peaks.</p> <p>Draw this structure.</p>	Was there a typo though? Regarding what? No dont think so, it was just because 2 of the peaks were the same environment almost so people got thrown off	5 2 1 2

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 - The Rise of Digital Reading Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Advantages of eBooks Over Traditional Books
2. Identifying Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unofficial Mark Scheme Chemistry 2013 Unit 2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Personalized Recommendations
 - Unofficial Mark Scheme Chemistry 2013 Unit 2 User Reviews and Ratings
 - Unofficial Mark Scheme Chemistry 2013 Unit 2 and Bestseller Lists
5. Accessing Unofficial Mark Scheme Chemistry 2013 Unit 2 Free and Paid eBooks

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 - Unofficial Mark Scheme Chemistry 2013 Unit 2 eBook Subscription Services
 - Unofficial Mark Scheme Chemistry 2013 Unit 2 Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Unofficial Mark Scheme Chemistry 2013 Unit 2 Compatibility with Devices
 - Unofficial Mark Scheme Chemistry 2013 Unit 2 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Highlighting and Note-Taking Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Interactive Elements Unofficial Mark Scheme Chemistry 2013 Unit 2
 8. Staying Engaged with Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unofficial Mark Scheme Chemistry 2013 Unit 2
 9. Balancing eBooks and Physical Books Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unofficial Mark Scheme Chemistry 2013 Unit 2
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Setting Reading Goals Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Fact-Checking eBook Content of Unofficial Mark Scheme Chemistry 2013 Unit 2
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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