

## SALT ANALYSIS

### TEST FOR ANIONS

EXPERIMENT	OBSERVATION	INFERENCE
1. Test for $\text{CO}_3^{2-}$ → A little of the salt is treated with dil. HCl.	Brisk effervescence with the evolution of colourless odourless gas	May be $\text{CO}_3^{2-}$ $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
Confirmatory test :- 2. Pass the above gas through lime water.	Lime water turns milky	Presence of $\text{CO}_3^{2-}$ is confirmed $\text{Ca(OH)}_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$
Test for Halides 3. A little of the salt is treated with 1 or 2 drops of conc. $\text{H}_2\text{SO}_4$	a) Colourless pungent smelling gas evolves, dense white fumes when a glass rod is dipped in $\text{NH}_4\text{OH}$ is shown on the mouth of the test tube b) Reddish brown vapours with pungent smell  c) Deep violet vapours with pungent smell	a) May be $\text{Cl}^-$ $\text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{NaHSO}_4 + \text{HCl}$  b) Maybe $\text{Br}^-$ $\text{NaBr} + \text{H}_2\text{SO}_4 \rightarrow \text{NaHSO}_4 + \text{HBr}$ $\text{H}_2\text{SO}_4 + 2\text{HBr} \rightarrow \text{SO}_2 + \text{Br}_2 + 2\text{H}_2\text{O}$  c) May be $\text{I}^-$ $\text{KI} + \text{H}_2\text{SO}_4 \rightarrow \text{KHSO}_4 + \text{HI}$ $\text{H}_2\text{SO}_4 + 2\text{HI} \rightarrow \text{SO}_2 + \text{I}_2 + 2\text{H}_2\text{O}$
Confirmatory test of halides 4. A little of the salt solution is acidified with dil. $\text{HNO}_3$ , boil, cool and add $\text{AgNO}_3$ sol.	a) Curdy white ppt. soluble in $\text{NH}_4\text{OH}$ .  b) Pale yellow ppt. sparingly soluble in $\text{NH}_4\text{OH}$  c) Yellow ppt. insoluble in $\text{NH}_4\text{OH}$	a) presence of $\text{Cl}^-$ is confirmed $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl (white ppt.)}$  b) presence of $\text{Br}^-$ is confirmed $\text{KBr} + \text{AgNO}_3 \rightarrow \text{KNO}_3 + \text{AgBr (pale yellow ppt.)}$  c) presence of $\text{I}^-$ is confirmed $\text{KI} + \text{AgNO}_3 \rightarrow \text{KNO}_3 + \text{AgI (yellow ppt.)}$

# Unknown Salt Identification

**Wilhelm Segerblom**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name bar.

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**College Practical Chemistry** V K Ahluwalia, Sunita Dhingra, Sunita Dhingra, 2005, **Salts and Their Reactions** Leonard Dobbin, Hugh Marshall, 1904 **Advanced Inorganic Analysis - I** Mr. Ajeet Singh, 2018-05-22 This book provides in depth coverage of modern methods in inorganic analysis including advanced qualitative and quantitative techniques instrumentation and interpretation of results catering to postgraduate and research level chemistry students **Abiotic stress: molecular genetics and genomics, Volume II** Mukesh Jain, Rohini Garg, Rajeev K. Varshney, Prasanta Kumar Subudhi, *EduGorilla's CBSE Class 11th Chemistry Lab Manual | 2024 Edition | A Well Illustrated, Complete Lab Activity book with Separate FAQs for Viva Voce Examination*, Need an informative and well illustrated Lab Manual CBSE Class 11th Chemistry Lab Manual is here for you The Lab Manual provides comprehensive steps for guiding students through each experiment Rigorously researched content prepared by a team of educators writers editors and proofreaders CBSE Class XI Chemistry Lab Manual has properly labeled high resolution diagrams and graphs A separate section on Viva Questions has been included to aid students in their Viva examination The Lab Manual explains the complex topics through detailed illustrations and lucid language making them simple to grasp Worksheets have been provided in CBSE Class 11th Chemistry Lab Manual for doing rough work Laboratory Manual for Chemistry Fundamentals Phyllis Buell, James Girard, 2002-08 Physical Sciences *Edexcel A Level Chemistry Student Book 1* Andrew Hunt, Graham Curtis, Graham Hill, 2015-07-17 Exam Board Edexcel Level AS A level Subject Chemistry First Teaching September 2015 First Exam June 2016 Endorsed by Edexcel Develop and assess your students knowledge and mathematical skills throughout A Level with worked examples practical assessment guidance and differentiated end of topic questions with this Edexcel Year 1 student book Identifies the level of your students understanding with diagnostic questions and a summary of prior knowledge at the start of the Year 1 Student Book Provides support for all 16 required practicals with various activities and questions along with a Practical chapter covering procedural understanding and key ideas related to measurement Mathematical skills are integrated throughout with plenty of worked examples including notes on methods to help explain the strategies for solving each type of problem Offers plenty of practice with Test Yourself Questions to help students assess their understanding and measure progress Encourages further reading and study with short passages of extension material Develops understanding with free online access to Test yourself Answers and an Extended Glossary Edexcel A level Chemistry Year 1 Student Book includes AS level Experiments in General Chemistry Kenneth W. Whitten, Kenneth D. Gailey, 1981 *General, Organic, and Biochemistry Lab Manual* Sara Selfe, 2006-01-12 Offers a choice of classic chemistry experiments and innovative ones All of them place special emphasis on the biological implications of chemical concepts Available for custom publishing at <http://custompub.whfreeman.com> *Lab Manual for Investigating Chemistry* David Collins (Ph. D.), Matthew Johll, 2008-12-02 While many of the core labs from the first edition have been retained a renewed focus on the basics of chemistry and the

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Turner,2020-11-03 A quirky oddly touching book that allows us to step for a few moments inside the world of a practicing Enlightenment scientist to sit beside him as he fans the flames of a candle with his little blowpipe waiting for that small mineral in front of him to melt and yield its secrets Wall Street Journal An accessible exploration of the noteworthy scientific career of James Smithson who left his fortune to establish the Smithsonian Institution James Smithson is best known as the founder of the Smithsonian Institution but few people know his full and fascinating story He was a widely respected chemist and mineralogist and a member of the Royal Society but in 1865 his letters collection of 10 000 minerals and more than 200 unpublished papers were lost to a fire in the Smithsonian Castle His scientific legacy was further written off as insignificant in an 1879 essay published through the Smithsonian fifty years after his death a claim that author Steven Turner demonstrates is far from the truth By providing scientific and intellectual context to his work The Science of James Smithson is a comprehensive tribute to Smithson s contributions to his fields including chemistry mineralogy and more This detailed narrative illuminates Smithson and his quest for knowledge at a time when chemists still debated thing as basic as the nature of fire and struggled to maintain their networks amid the ever changing conditions of the French Revolution and the Napoleonic Wars

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