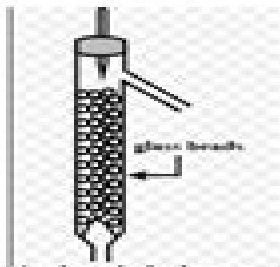


Suggest why chloroethanoic acid is stronger than ethanoic acid	(this question is already discussed a little further down.)		
fill in rate table	1.00 0.16 1.35 I beg someone write what the values were if they remember	K=250 Yes i agree - needs to be 1.4 as all values were to two sig figs Weren't we they to 3 dp though? The rate of the first one was 1.00 so I'm guessing 3sf of 2dp	3
empirical formula	C6H10O3S2	Percentages dudbt add up to 100 fuck me Didn't even see the sulphur, fml same same i thought i meant i was in the fuel area. YAY i got this!!	3
the labeling distillation	for the tube it was fractionating column or glass beads the bottom of the condenser was water in and top was water out	 reject anti bumping granules might accept silica BEADS	3
Draw a line for the reaction with chloroethanoic acid Suggest why chloroethanoic acid is a stronger acid than ethanoic acid.	Was the line of chloroethanoic acid to the left of the first line, but levelled off at the same point? Yes but it levelled out before, same height though. Ah ok. Dissociates more Chlorine is more electronegative than hydrogen so will withdraw electron density from the carboxyl group meaning that the O-H bond breaks more easily dissociating into H⁺ making the solution more acidic	Cl is electron withdrawing. it polarises the bond, weakens the O-H bond which releases more H ⁺ than ethanoic acid Neg inductive effect? Chloroethanoic acid forms a more stable conjugate base?	3
NMR Question?	Use the integration values and the ratios to identify the structure at two peaks. Draw this structure.	Was there a typo though? Regarding what? No dont think so, it was just because 2 of the peaks were the same environment almost so people got thrown off	5 2 1 2

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Table of Contents Unofficial 2015 4h Mark Scheme

1. Understanding the eBook Unofficial 2015 4h Mark Scheme
 - The Rise of Digital Reading Unofficial 2015 4h Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying Unofficial 2015 4h Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unofficial 2015 4h Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unofficial 2015 4h Mark Scheme
 - Personalized Recommendations
 - Unofficial 2015 4h Mark Scheme User Reviews and Ratings
 - Unofficial 2015 4h Mark Scheme and Bestseller Lists
5. Accessing Unofficial 2015 4h Mark Scheme Free and Paid eBooks
 - Unofficial 2015 4h Mark Scheme Public Domain eBooks
 - Unofficial 2015 4h Mark Scheme eBook Subscription Services
 - Unofficial 2015 4h Mark Scheme Budget-Friendly Options
6. Navigating Unofficial 2015 4h Mark Scheme eBook Formats

- ePub, PDF, MOBI, and More
- Unofficial 2015 4h Mark Scheme Compatibility with Devices
- Unofficial 2015 4h Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unofficial 2015 4h Mark Scheme
 - Highlighting and Note-Taking Unofficial 2015 4h Mark Scheme
 - Interactive Elements Unofficial 2015 4h Mark Scheme
- 8. Staying Engaged with Unofficial 2015 4h Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unofficial 2015 4h Mark Scheme
- 9. Balancing eBooks and Physical Books Unofficial 2015 4h Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unofficial 2015 4h Mark Scheme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unofficial 2015 4h Mark Scheme
 - Setting Reading Goals Unofficial 2015 4h Mark Scheme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unofficial 2015 4h Mark Scheme
 - Fact-Checking eBook Content of Unofficial 2015 4h Mark Scheme
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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