Name \_\_\_\_\_

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## Chemistry - Unit 6 Worksheet 2

Balance the following equations by inserting the proper coefficients.

- 1. \_\_\_\_SO<sub>2</sub> + \_\_\_\_SO<sub>3</sub>
- 2. \_\_\_CH<sub>4</sub> + \_\_\_O<sub>2</sub> ---> \_\_\_CO + \_\_\_H<sub>2</sub>O
- 3. \_\_\_P + \_\_\_Cl2 ---> \_\_\_PCl3
- 4. \_\_\_CO + \_\_\_O<sub>2</sub> ---> \_\_\_CO<sub>2</sub>
- CH<sub>4</sub> + \_\_\_ O<sub>2</sub> ---> \_\_ CH<sub>3</sub>OH
- 6. \_\_\_\_Li + \_\_\_\_Br2 ---> \_\_\_LiBr
- 7. \_\_\_\_Al<sub>2</sub>O<sub>3</sub> ---> \_\_\_\_Al + \_\_\_O<sub>2</sub>
- Na + \_\_\_\_H<sub>2</sub>O ---> \_\_\_\_NaOH + \_\_\_H<sub>2</sub>
- 9. \_\_\_CO2 + \_\_\_H2O ---> \_\_\_C6H12O6 + \_\_\_O2
- 10. H2SO4 + NaCl ---> HCl + Na2SO4

Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:

- 11. Hp + SOp ---> HpS + HpO
- 12. \_\_\_CaCO<sub>3</sub> + \_\_\_SO<sub>2</sub> + \_\_\_O<sub>2</sub> ---> \_\_\_CaSO<sub>4</sub> + \_\_\_CO<sub>2</sub>
- \_\_\_AgNO<sub>3</sub> + \_\_\_CaCl<sub>2</sub> ---> \_\_\_AgCl + \_\_\_Ca(NO<sub>3</sub>)<sub>2</sub>
- 14. \_\_\_\_HCl + \_\_\_\_Ba(OH)2 ---> \_\_\_\_BaCl2 + \_\_\_\_H2O
- 15. \_\_\_\_H3PO4 + \_\_\_\_NaOH ---> \_\_\_\_Na3PO4 + \_\_\_\_H2O
- 16. \_\_\_\_Pb(NO3)2 + \_\_\_\_KI ---> \_\_\_\_Pbl2 + \_\_\_KNO3
- 17. \_\_\_CuO + \_\_\_NH<sub>3</sub> ---> \_\_\_N<sub>2</sub> + \_\_\_Cu + \_\_\_H<sub>2</sub>O
- 18. \_\_\_\_C<sub>2</sub>H<sub>5</sub>OH + \_\_\_\_O<sub>2</sub> ---> \_\_\_\_CO<sub>2</sub> + \_\_\_\_H<sub>2</sub>O
- 19. \_\_\_\_C2H6 + \_\_\_\_O2 ---> \_\_\_\_CH3COOH + \_\_\_\_H2O
- 20. \_\_\_NO<sub>2</sub> + \_\_\_H<sub>2</sub>O ---> \_\_\_HNO<sub>3</sub> + \_\_\_NO

Write the formulas of the reactants and products - including the symbols for the state - (s), (l), (g), (aq) - then balance the equations.

 Gaseous ammonia (NH<sub>3</sub>) reacts with hydrogen chloride gas to form solid ammonium chloride.

# **Tesccc Unit 9 Lesson 1 Chemistry Answers**

**G Psacharopoulos** 

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