

**MODIFIED TRUE/FALSE**

1. ANS: F, an increase
2. ANS: F, bottom right side
3. ANS: T
4. ANS: F, negative
5. ANS: F, hydrogen gas at the cathode and oxygen gas at the anode

**MULTIPLE CHOICE**

6. ANS: D      7. ANS: B      8. ANS: C      9. ANS: A      10. ANS: B

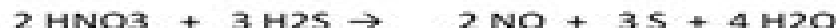
**11.** Define oxidation number, oxidation, reduction, reducing agent and oxidizing agent.  
Refer to your notes.

**12.** What is oxidized and what is reduced in the following reaction:  $\text{N}_2 + 2\text{O}_2 \rightarrow 2\text{NO}_2$   
 $\text{N}_2$  is oxidized;  $\text{O}_2$  is reduced

**13.** Determine the oxidation number of each element in the following substances:

- a )  $\text{H}_2\text{SO}_4$  H: +1, S: +6, O: -2    b )  $\text{Ca}_3(\text{PO}_4)_2$  Ca: +2, P: +5, O: -2    c )  $\text{Cl}_2$  Cl: 0  
d )  $\text{KMnO}_4$  K: +1, Mn: +7, O: -2      e )  $\text{H}_3\text{AsO}_4$  H: +1, As: +5, O: -2      f )  $\text{CaSnO}_3$  Ca: +2, Sn: +4, O: -2

**14.** Balance the following equations (use the oxidation number method):



**15.** Iodine normally has oxidation numbers of -1 and 0. Could  $\text{I}_2$  serve as a reducing agent? Explain.  
Answer: This is not likely to happen because for  $\text{I}_2$  to be oxidized it would have to lose an electron forming a +1 oxidation number which iodine does not typically do.

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Test Electrochemistry Answers**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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