

Assessment

Chapter Test B**Chapter: Gases**

PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. If the temperature of a gas remains constant, then the pressure of the gas will increase if the
- mass of the gas molecules decreases.
 - diffusion of the gas molecules increases.
 - size of the container is decreased.
 - number of gas molecules in the container is decreased.
- _____ 2. When Gay-Lussac's law of combining volumes holds, which of the following can be expressed in ratios of small whole numbers?
- pressures before and after reaction
 - volumes of gaseous reactants and products
 - kelvin temperatures
 - molar masses of products and molar masses of reactants
- _____ 3. Equal volumes of ideal gases at the same temperature and pressure contain equal numbers of
- protons.
 - ions.
 - particles.
 - electrons.
- _____ 4. At constant temperature and pressure, the volume of a gas is directly proportional to its
- molar mass.
 - number of moles.
 - density at STP.
 - rate of diffusion.
- _____ 5. To use the ideal gas law to determine the molar mass of a gas, you
- must determine the mass of a molar volume of that gas.
 - may use the mass of any known volume of the gas.
 - may not use a volume of less than 22.4 L.
 - must make the volume measurement at STP.
- _____ 6. Suppose that two unlike gases are injected into opposite ends of a long tube at the same time and allowed to diffuse toward the center. They should begin to mix
- at the end that held the heavier gas.
 - between the end where the heavier gas entered and the middle.
 - between the end where the lighter gas entered and the middle.
 - exactly in the middle.

Test Rinehart And Winston

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