

$$\lambda = \frac{3.00 \times 10^8 \text{ m s}^{-1}}{2.34 \times 10^{14} \text{ s}^{-1}}$$

$$= 1.28 \times 10^{-6} \text{ m} = 1.3 \mu\text{m}$$

The answer is C.

- b) The red end of the visible is about 700 nm ($= 0.7 \mu\text{m}$), so $1.3 \mu\text{m}$ will be just beyond the red, i.e. in the infrared.

The answer is A.

- 4 a) A photon is a small, discrete amount, or quantum, of energy associated with electromagnetic radiation.

- b) The energy to move an electron through a potential difference of 0.48 V is given by:

$$E = QV = 1.6 \times 10^{-19} \text{ C} \times 0.48 \text{ V} = 7.68 \times 10^{-20} \text{ J} \approx 8 \times 10^{-20} \text{ J}$$

- c) The efficiency of energy conversion is therefore:

$$\text{efficiency} = \frac{7.68 \times 10^{-20} \text{ J}}{4 \times 10^{-19} \text{ J}} \times 100\% = 19\%$$

- d) Rearranging $E = hf$:

$$f = \frac{E}{h} = \frac{4.0 \times 10^{-19} \text{ J}}{6.63 \times 10^{-34} \text{ J s}} = 6.03 \times 10^{14} \text{ s}^{-1}$$

Rearranging $c = f\lambda$:

$$\lambda = \frac{c}{f} = \frac{3.0 \times 10^8 \text{ m s}^{-1}}{6.03 \times 10^{14} \text{ s}^{-1}} = 4.97 \times 10^{-7} \text{ m} = 500 \text{ nm (to 2 significant figures)}$$

- 5 a) The work function of a surface is the amount of energy that is needed to just remove an electron from the surface.

- b) We will need to use $E = hf$, so we must find the frequency corresponding to a wavelength of 532 nm. Rearranging $c = f\lambda$:

$$f = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m s}^{-1}}{532 \times 10^{-9} \text{ m}} = 5.64 \times 10^{14} \text{ s}^{-1}$$

Then:

$$E = hf = 6.63 \times 10^{-34} \text{ J s} \times 5.64 \times 10^{14} \text{ s}^{-1} = 3.74 \times 10^{-19} \text{ J}$$

Converting to electron-volts using $1 \text{ eV} = 1.6 \times 10^{-19} \text{ J}$:

$$E = \frac{3.74 \times 10^{-19} \text{ J}}{1.6 \times 10^{-19} \text{ J eV}^{-1}} = 2.34 \text{ eV} \approx 2.3 \text{ eV}$$

- c) Rearranging $hf = \phi + \frac{1}{2}mv_{\text{max}}^2$ gives:

$$\frac{1}{2}mv_{\text{max}}^2 = hf - \phi = (2.34 - 1.90) \text{ eV} = 0.44 \text{ eV}$$

$$= 0.44 \text{ eV} \times 1.6 \times 10^{-19} \text{ J eV}^{-1} = 7.0 \times 10^{-20} \text{ J}$$

- d) The frequency of the red light will be less than that of the green light, so a quantum of the red light may not have sufficient energy (hf) to overcome the work function and release photoelectrons.

A quick way of calculating the wavelength corresponding to the work function is to say that 2.34 eV corresponds to 532 nm (from part b)), therefore 1.90 eV (the work function) will correspond to:

$$\lambda = \left(\frac{2.34}{1.90} \right) \times 532 \text{ nm} = 655 \text{ nm}$$

This is in the red region of the spectrum and so if the laser has a wavelength greater than this, it will not cause any photoelectrons to be emitted.

- 6 a) Green light has the shortest wavelength and therefore the highest frequency. The green LED will therefore emit photons of the highest energy as the photon energy is given by $E = hf$.

- b) i) We will need to use $E = hf$, so we must find the frequency corresponding to a wavelength of 630 nm. Rearranging $c = f\lambda$:

Tip

If you use the "show that" value of $8 \times 10^{-20} \text{ J}$ and get 20% you will still get full marks, but it is better to use the actual calculated value.

Tip

As the question asks you to suggest the reason, you would not be expected to go into as much detail as this – the first paragraph would be sufficient.

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