CHAPTER 19 ELECTROCHEMISTRY

Problem Categories

Biological: 19.66, 19.68, 19.126.

Conceptual: 19.67, 19.79, 19.85, 19.94, 19.103, 19.106, 19.108, 19.128.

Descriptive: 19.13, 19.14, 19.17, 19.18, 19.46a, 19.52a, 19.61, 19.77, 19.87, 19.93, 19.99, 19.100, 19.101, 19.111, 19.113, 19.123.

Environmental: 19.63.

Budistrial: 19.48, 19.56, 19.89, 19.109, 19.112, 19.120.

Organic: 19.38.

Difficulty Level

Easy: 19.11, 19.12, 19.13, 19.14, 19.16, 19.17, 19.18, 19.22, 19.45, 19.63, 19.67, 19.79, 19.91, 19.97,
**Mechanical Societies of the Computation of the Computation

- 19.1 We follow the steps are described in detail in Section 19.1 of the text.
 - (a) The problem is given in ionic form, so combining Steps 1 and 2, the half-reactions are:

oxidation: Fe²⁺ → Fe³⁺
reduction: H₂O₂ → H₂O

Step 3: We balance each half-reaction for number and type of atoms and charges.

The oxidation half-reaction is already balanced for Fe atoms. There are three net positive charges on the right and two net positive charges on the left, we add one electrons to the right side to balance the charge.

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$

Reduction half-reaction: we add one H₂O to the right-hand side of the equation to balance the O atoms.

$$H_2O_2 \rightarrow 2H_2O$$

To balance the H atoms, we add 2H" to the left-hand side.

$$H_2O_2 + 2H^* \rightarrow 2H_2O$$

There are two net positive charges on the left, so we add two electrons to the same side to balance the charge.

$$H_2O_2 + 2H'' + 2e'' \rightarrow 2H_2O$$

Step 4: We now add the exidation and reduction half-reactions to give the overall reaction. In order to equalize the number of electrons, we need to multiply the exidation half-reaction by 2.

$$2(Fe^{2+} \rightarrow Fe^{3+} + e^{-})$$

 $H_2O_2 + 2H^{-} + 2e^{-} \rightarrow 2H_2O$
 $2Fe^{2+} + H_2O_2 + 2H^{-} + 2e^{-} \rightarrow 2Fe^{3+} + 2H_2O + 2e^{-}$

Test Exam Of Electrochemistry Chap 19

Rachel S Tattersall

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