## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- 1. The charge on an ion is
  - a. always positive.
  - b. always negative.
  - c. either positive or negative.
  - d. zero.
- According to the octet rule, a calcium atom has a tendency to
  - a. lose one electron.
  - b. lose two electrons.
  - c. gain one electron.
  - d. gain two electrons.
- 3. If a compound forms by ionic bonding, which is not true?
  - a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
  - The net charge of the compound is zero.
  - c. The compound contains just two atoms, each of opposite charge.
  - Several ions group together in a tightly packed structure.
- 4. The only property listed that is not characteristic of ionic compounds is
  - a. high melting point.
  - b. hardness.
  - c. lack of crystal structure.
  - d. good conductor of electricity.
- 5. Which formula listed below represents a polyatomic ion?
  - a. NH<sub>4</sub>\*
  - b. H<sub>2</sub>SO<sub>4</sub>
  - c. NaCl.
  - d. H<sub>2</sub>O
- 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
  - a. ionic substances tend to vaporize at room temperature.
  - b. ionic substances are brittle.
  - attractive forces between ions are greater than the attractive forces between molecules.
  - d. the numbers of positive and negative charges are equal in an ionic compound.

# **Test Chemical Bonding Answer Key**

John T. Moore, Richard H. Langley

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