

MODIFIED TRUE/FALSE

1. ANS: F, an increase
2. ANS: F, bottom right side
3. ANS: T
4. ANS: F, negative
5. ANS: F, hydrogen gas at the cathode and oxygen gas at the anode

MULTIPLE CHOICE

6. ANS: D 7. ANS: B 8. ANS: C 9. ANS: A 10. ANS: B

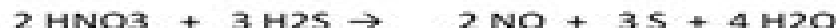
11. Define oxidation number, oxidation, reduction, reducing agent and oxidizing agent.
Refer to your notes.

12. What is oxidized and what is reduced in the following reaction: $\text{N}_2 + 2\text{O}_2 \rightarrow 2\text{NO}_2$
 N_2 is oxidized; O_2 is reduced

13. Determine the oxidation number of each element in the following substances:

- a) H_2SO_4 H: +1, S: +6, O: -2 b) $\text{Ca}_3(\text{PO}_4)_2$ Ca: +2, P: +5, O: -2 c) Cl_2 Cl: 0
d) KMnO_4 K: +1, Mn: +7, O: -2 e) H_3AsO_4 H: +1, As: +5, O: -2 f) CaSnO_3 Ca: +2, Sn: +4, O: -2

14. Balance the following equations (use the oxidation number method):



15. Iodine normally has oxidation numbers of -1 and 0. Could I_2 serve as a reducing agent? Explain.
Answer: This is not likely to happen because for I_2 to be oxidized it would have to lose an electron forming a +1 oxidation number which iodine does not typically do.

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