

# Stoichiometry "End of Chapter" IN CLASS Practice Assessment

## Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?  
Coefficients are necessary for establishing a molar/mole ratio to convert between moles of different substances.

**Skill 1** - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

**Skill 2** - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[ \frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[ \frac{2.8\text{g}/4.0\text{g}}{1} \right] \times 100 = 70\%$$

**Skill 3** - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol CaCO}_3 \times \frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} = 1.55\text{mol CaCl}_2$$

$$250\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} = 6.857\text{mol HCl} \quad \text{E.R.} \\ 6.857\text{mol HCl} \times \frac{1\text{mol CaCl}_2}{2\text{mol HCl}} = 3.43\text{mol CaCl}_2$$

$$1.55\text{mol CaCO}_3 \times \left( \frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left( \frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[ \frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[ \frac{142.0\text{g}/172.02\text{g}}{1} \right] \times 100 = 82.55\%$$

# Test Review Stoichiometry Answers

**The Princeton Review**



## Test Review Stoichiometry Answers:

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## Reviewing **Test Review Stoichiometry Answers**: Unlocking the Spellbinding Force of Linguistics

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### **Table of Contents Test Review Stoichiometry Answers**

1. Understanding the eBook Test Review Stoichiometry Answers
  - The Rise of Digital Reading Test Review Stoichiometry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Test Review Stoichiometry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Test Review Stoichiometry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Test Review Stoichiometry Answers
  - Personalized Recommendations
  - Test Review Stoichiometry Answers User Reviews and Ratings
  - Test Review Stoichiometry Answers and Bestseller Lists

5. Accessing Test Review Stoichiometry Answers Free and Paid eBooks
  - Test Review Stoichiometry Answers Public Domain eBooks
  - Test Review Stoichiometry Answers eBook Subscription Services
  - Test Review Stoichiometry Answers Budget-Friendly Options
6. Navigating Test Review Stoichiometry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Test Review Stoichiometry Answers Compatibility with Devices
  - Test Review Stoichiometry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Test Review Stoichiometry Answers
  - Highlighting and Note-Taking Test Review Stoichiometry Answers
  - Interactive Elements Test Review Stoichiometry Answers
8. Staying Engaged with Test Review Stoichiometry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Test Review Stoichiometry Answers
9. Balancing eBooks and Physical Books Test Review Stoichiometry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Test Review Stoichiometry Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Test Review Stoichiometry Answers
  - Setting Reading Goals Test Review Stoichiometry Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Test Review Stoichiometry Answers
  - Fact-Checking eBook Content of Test Review Stoichiometry Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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