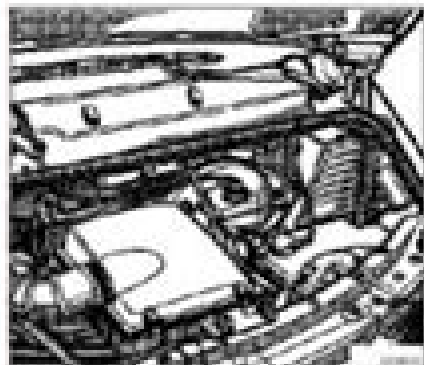


Vauxhall Movano 1998-2003 Electrical Systems



A - CHECKING

Check and ensure that:

- the battery tray and cover are not cracked or split.
- the top of the battery is clean.
- the terminals are in good condition.

B - PRECAUTIONS

It should be remembered that a battery:

- contains sulphuric acid, which is a dangerous product.
- produces oxygen and hydrogen during charging. The mixture of these two gases forms a detonating gas, hence the risk of explosion.

C - DANGER - ACID

The sulphuric acid which is contained in the battery is extremely toxic and corrosive. It can burn the skin, damage clothing, corrode the terminals.

It is also very important when handling a battery, to take the following precautions:

- to protect your eyes with goggles.
- to wear acid gloves and clothing.

It is vital:

- to ensure that there is no sulphation on the terminals.
- to clean and grease the terminals if necessary.
- to check that the nuts are correctly tightened on the terminals. Loose nuts could cause starting faults or charging faults which could cause sparks, making the battery liable to explode.
- to check the electrolyte level.

Batteries with sets of removable plugs:

- remove the cover by hand or by using a screw drill operator.
- check that the electrolyte level in all the cells is well above the level of the separators.
- if necessary, use deionised/distilled water to top up the level.

Note: certain types of battery have venting facilities which allow the build-up of gas to be vented.

Note: add electrolyte to the cells nearest the battery.

D - DANGER - HYDROGEN

When a battery is being charged in a vehicle or when the battery is being charged in a workshop, hydrogen gas is produced. This gas is highly explosive and the quantity of gas produced is proportional to the intensity of the charging current.

The oxygen and the hydrogen join together in the open air, on the surface of the plates and form a highly explosive mixture.

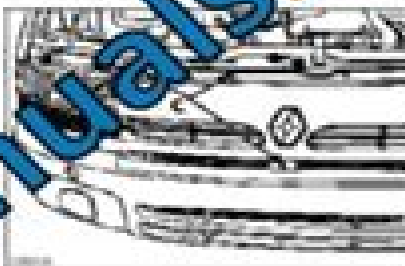
The smallest of sparks, a cigarette or a recently extinguished match are sufficient to cause an explosion. The explosion is so strong that the battery can shatter and the acid is dispersed into the surrounding atmosphere. People nearby are at risk (blowback, scalding, pain, acid splashes). The acid splashes are harmful to the eyes, face and hands. They also attack clothing.

REMOVAL - BATTERY

Disconnect:

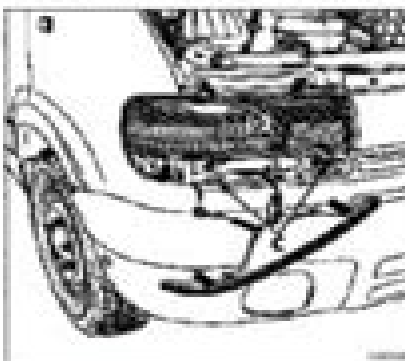
- the battery.
- the connectors for the bulbs in the bonnet (the indicator and for the rear adjustment unit).

Remove the three screws (1), (2) and (3) from the radiator grille.



Remove:

- the indicator.
- the two mounting bolts (2) from the radiator grille.
- the four mounting bolts (2) from the bonnet.



The headlight is retained in the engine room to prevent it from falling.

REMOVAL - BATTERY

Disconnect the battery from the engine.



Turn the engine on slightly of a turn towards the left to bring it to release the bonnet.

Then release the ball joint of the engine and the bonnet parallel to remove the engine.

SPECIAL NOTES FOR REMOVAL

Remove the bulb connector protective cover at the rear of the bonnet.

Keep the parallel towards the rear of the bonnet by pulling on the base of the bulb, and check the ball joint into position.

Now, roll the engine to the bonnet by turning it on slightly of a turn.

Reconnect the connector and roll the protective bulb cover at the rear of the bonnet.

Set the engine adjustment control to 'V' and adjust the headlight.

- ball (A) for height adjustment.
- ball (B) for direction adjustment.

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web this volume is known as the molar gas volume V_m the volume of a gas may be calculated from its number of moles using volume of gas moles V_m rearranging this moles of gas $\frac{V}{V_m}$

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web jun 8 2023 $V = 1.00 \text{ mol} \times 0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1} \times 273 \text{ K} = 22.4 \text{ L}$ in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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