Stoichiometry Study Guide KEY

Chemistry

RHS - Mr. Moss

- Define the following:
 - a. Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction.
 - b. Mole The SI unit used to measure the amount of a substance that contains 6.02 x 10²³ atoms of that substance.
 - Mole Ratio-The ratio between any two substances in a balanced chemical equation.
 - d. Molar Mass-The mass in grams of one mole of any substance.
- What relationships can be determined from a balanced chemical equation? The relationships between particles, moles, and mass for all reactants and products.
- Explain why mole ratios are central to stoichiometric calculations. Mole ratios allow for the conversion from moles of one substance in a balanced chemical equation to moles of another substance in the same equation.
- What is the first step in all stoichiometric calculation? (Be careful.) Write a balanced chemical equation.
- How is molar mass used in some stoichiometric calculations? Molar mass is a conversion factor for converting moles of a given substance to mass or mass of a given substance to moles.
- 6. Find the molar mass of the following:
 - a. H2SO4 (sulfuric acid) 98.09 g/mol
 - b. ZnO (zinc oxide) 81.39 g/mol
 - NH₄NO₃ (ammonium nitrate) 80.05 g/mol
- 7. List the mole ratios of the following balanced chemical equation.
 - a. 2Al + 3Br₂ → 2AlBr₃ 2 mol Al 3 mol Br₂ ; 3 mol Br₂ ; 2 mol AlBr₃ ; 2 mol AlBr₃ (Each one of these can be inverted to form another mole ratio. This would give a total of 6 ratios.)

Study Guide Stoichiometry Key

William H. Myers, William T. Mooney

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