

# Acids and Bases

Definitions	Acid	Base
1. Bronsted-Lowry	H <sup>+</sup> donor	H <sup>+</sup> acceptor
2. Lewis	e <sup>-</sup> acceptor	e <sup>-</sup> donor

## Typical Acid/Base Reaction:



**K<sub>eq</sub> for this example:** **K<sub>a</sub> for this example:**

$$K_{eq} = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{H}_2\text{O}][\text{HA}]} \quad K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$

## In general:

- Strong acid → Weak conj. base
- Weak acid → Strong conj. base
- Strong base → Weak conj. acid
- Weak base → Strong conj. acid

## Predicting equilibria:

Start by labeling the pK<sub>a</sub>s in the reaction (see list)

Reaction Arrows:

- R<sub>2</sub>N arrows point towards the side with the weaker acid (higher pK<sub>a</sub>)

Is it reversible or irreversible?

- Look at the difference of pK<sub>a</sub>s between reactants/products
- $\Delta \text{pK}_a > 8 \Rightarrow$  irreversible rxn
- $\Delta \text{pK}_a < 8 \Rightarrow$  reversible rxn

Example: Predict equilibrium using pK<sub>a</sub>s, is it reversible or irreversible?



Acetylene **pK<sub>a</sub> 25**

1. Label pK<sub>a</sub>s
2. Know arrows will point towards weaker acid (higher pK<sub>a</sub>), so left
3. Is it reversible or irreversible? (25-16 > 8), so irreversible
4. Final arrows:  $\rightarrow$

Things to know about acid/base reactions:

**Equilibrium constant (K<sub>eq</sub>):** a ratio that tells the position of a chemical equilibrium

$$K_{eq} = \frac{[\text{products}]}{[\text{reactants}]}$$

Trends:

- K<sub>eq</sub> > 1  $\Rightarrow$  Favors products
- K<sub>eq</sub> = 1  $\Rightarrow$  At equilibrium
- K<sub>eq</sub> < 1  $\Rightarrow$  Favors reactants (non doesn't work)

**Acid Dissociation Constant (K<sub>a</sub>):** measure of acid strength

$$K_a = K_{eq} [\text{H}_2\text{O}]$$

Trends:

- $\uparrow K_a =$ 
  - Strong acid
  - Full dissociation
  - High [products] (large numerator)
- $\downarrow K_a =$ 
  - Weak acid
  - Partial/incomplete dissociation
  - High [reactants] (large denominator)

$$\text{pK}_a = -\log K_a$$

Trends:

- $\downarrow \text{pK}_a =$  Strong acid  $\uparrow \text{pK}_a =$  Weak acid
- $\uparrow \text{pK}_a =$  Weak acid  $\downarrow \text{pK}_a =$ 
  - pK<sub>a</sub> < 1  $\Rightarrow$  Strong acid
  - pK<sub>a</sub> 1-3  $\Rightarrow$  Moderately strong acid
  - pK<sub>a</sub> 3-5  $\Rightarrow$  Weak acid
  - pK<sub>a</sub> > 5  $\Rightarrow$  Very weak acid

## pK<sub>a</sub>s to Know

- sp<sup>3</sup> C-H 51
- sp<sup>2</sup> C-H 44
- sp C-H 25
- Water H<sub>2</sub>O = 16 (15.7)
- Ethanol OH = 16 (15.9)
- Ammonia NH<sub>3</sub> 38
- Acetic Acid CO<sub>2</sub>H = 5 (4.76)
- HCl -7

# Study Guides For Chemistry Acids And Bases

**Kevin Smith**



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## **Table of Contents Study Guides For Chemistry Acids And Bases**

1. Understanding the eBook Study Guides For Chemistry Acids And Bases
  - The Rise of Digital Reading Study Guides For Chemistry Acids And Bases
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guides For Chemistry Acids And Bases
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guides For Chemistry Acids And Bases
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guides For Chemistry Acids And Bases
  - Personalized Recommendations
  - Study Guides For Chemistry Acids And Bases User Reviews and Ratings
  - Study Guides For Chemistry Acids And Bases and Bestseller Lists
5. Accessing Study Guides For Chemistry Acids And Bases Free and Paid eBooks
  - Study Guides For Chemistry Acids And Bases Public Domain eBooks
  - Study Guides For Chemistry Acids And Bases eBook Subscription Services
  - Study Guides For Chemistry Acids And Bases Budget-Friendly Options

6. Navigating Study Guides For Chemistry Acids And Bases eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guides For Chemistry Acids And Bases Compatibility with Devices
  - Study Guides For Chemistry Acids And Bases Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guides For Chemistry Acids And Bases
  - Highlighting and Note-Taking Study Guides For Chemistry Acids And Bases
  - Interactive Elements Study Guides For Chemistry Acids And Bases
8. Staying Engaged with Study Guides For Chemistry Acids And Bases
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guides For Chemistry Acids And Bases
9. Balancing eBooks and Physical Books Study Guides For Chemistry Acids And Bases
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guides For Chemistry Acids And Bases
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guides For Chemistry Acids And Bases
  - Setting Reading Goals Study Guides For Chemistry Acids And Bases
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guides For Chemistry Acids And Bases
  - Fact-Checking eBook Content of Study Guides For Chemistry Acids And Bases
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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