

**Assignment:**  
**Index the Stoichiometric Fraction**

Name: \_\_\_\_\_

**Questions 1 – 3** are based on the following chemical reaction:



1. How many moles of  $\text{H}_2$  can be produced from 1.8 moles  $\text{Al}$ ?
2. If 3.6 moles of  $\text{H}_2\text{SO}_4$  reacts with excess  $\text{Al}$ , how many moles of  $\text{Al}_2(\text{SO}_4)_3$  will be produced?
3. How many moles of aluminum will be needed to produce 18.0 moles of hydrogen gas?

**Questions 4 – 6** are based on the decomposition of potassium carbonate. When heated, potassium carbonate decomposes into potassium oxide and oxygen.

**Balance Equation:** \_\_\_\_\_

4. How many moles of  $\text{O}_2$  can be produced from the decomposition of 1.00 moles  $\text{K}_2\text{CO}_3$ ?
5. If 6.2407 moles of  $\text{K}_2\text{CO}_3$  decomposes, how many moles of  $\text{K}_2\text{O}$  will be produced?
6. How many moles of potassium chloride will be needed to produce 18.0 moles of oxygen gas?

**Challenge:** How many moles of oxygen gas will be produced from the decomposition of 20.0 grams of  $\text{K}_2\text{CO}_3$ ?

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