

## FORMULAS FOR IONIC COMPOUNDS

1. Write the chemical formula for the ionic compounds made from the following elements.

Elements	metal ion	non-metal ion	Chemical Formula	Name
potassium and iodine	$K^{+1}$	$I^{-1}$	KI	Potassium Iodide
barium and chlorine	$Ba^{2+}$	$Cl^{-1}$	$BaCl_2$	Barium Chloride
sodium and oxygen	$Na^{+1}$	$O^{2-}$	$Na_2O$	sodium Oxide
aluminum and nitrogen	$Al^{3+}$	$N^{3-}$	AlN	Aluminum Nitride
lithium and sulfur	$Li^{+1}$	$S^{2-}$	$Li_2S$	Lithium Sulfide
magnesium and oxygen	$Mg^{+2}$	$O^{2-}$	MgO	Magnesium Oxide
calcium and chlorine	$Ca^{2+}$	$Cl^{-1}$	$CaCl_2$	calcium chloride
boron and chlorine	$B^{3+}$	$Cl^{-1}$	$BCl_3$	boron chloride
beryllium and sulfur	$Be^{2+}$	$S^{2-}$	BeS	beryllium sulfide

2. Complete the chart below and write the ionic formulas to complete it. A few have been done for you.

	F	Cl	$O^{2-}$	$S^{2-}$	$N^{3-}$	$P^{3-}$
$Li^{+}$	LiF	LiCl	$Li_2O$	$Li_2S$	$Li_3N$	$Li_3P$
$Na^{+}$	NaF	NaCl	Na <sub>2</sub> O	$Na_2S$	$Na_3N$	$Na_3P$
$K^{+}$	KF	KCl	$K_2O$	$K_2S$	$K_3N$	$K_3P$
$Be^{2+}$	$BeF_2$	$BeCl_2$	BeO	BeS	$Be_3S_2$	$Be_3P_2$
$Mg^{2+}$	$MgF_2$	$MgCl_2$	MgO	MgS	$Mg_3N_2$	$Mg_3P_2$
$Ca^{2+}$	$CaF_2$	$CaCl_2$	CaO	CaS	$Ca_3N_2$	$Ca_3P_2$
$Al^{3+}$	$AlF_3$	$AlCl_3$	$Al_2O_3$	$Al_2S_3$	AlN	AlP
$Ga^{3+}$	$GaF_3$	$GaCl_3$	$Ga_2O_3$	$Ga_2S_3$	GaN	GaP
$B^{3+}$	$BF_3$	$BCl_3$	$B_2O_3$	$B_2S_3$	BN	BP

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