

Study Buddy: Classification of Chemical Reactions

Terms	Definitions	Representation									
Law of Conservation of Mass	During a chemical or physical change, mass is conserved. Mass is neither created nor destroyed. Mass of reactants equals mass of products.	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ <table border="1"> <tr> <td>N_2</td><td>3H_2</td><td>$\rightarrow 2\text{NH}_3$</td></tr> <tr> <td>28g</td><td>6g</td><td>34g</td></tr> <tr> <td>2 atoms</td><td>6 atoms</td><td>8 atoms</td></tr> </table>	N_2	3H_2	$\rightarrow 2\text{NH}_3$	28g	6g	34g	2 atoms	6 atoms	8 atoms
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Precipitation Reaction	Occurs when two aqueous solutions react to produce a solid precipitate.	$\text{NaCl}_{(aq)} + \text{AgNO}_{3(aq)} \rightarrow \text{AgCl}_{(s)} + \text{NaNO}_{3(aq)}$									
Redox Reaction	Occurs when one or more electrons is transferred during a chemical reaction. Reduction: occurs when a species gains an electron Oxidation: occurs when a species loses an electron	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$ $\text{Mg} \rightarrow \text{Mg}^{2+} + 2\text{e}^-$									

Representations:

- Balance the following chemical equations and draw a particle diagram showing conservation of mass:

- Hydrogen gas and oxygen gas react to form liquid water.



- Solid zinc is added to an aqueous solution of silver nitrate to produce zinc nitrate and solid silver.



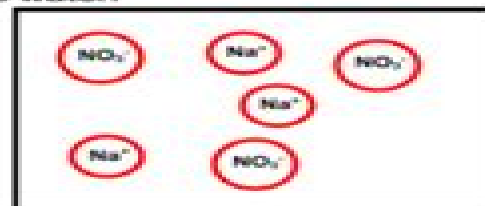
- A chemist added a 5.0g piece of solid magnesium metal to 10.0g of hydrochloric acid ($\text{HCl}_{(aq)}$) in a 50 mL beaker. During the reaction, the solution in the beaker became warmer and produced many bubbles. Once the reaction finished, the mass of the final solution was only 14.6g. What is the best explanation for the "missing" mass?

A bubbles are an indicator that a gas is being produced. In this case, 0.4g of gas were produced and no longer in the solution.

0.21 mole of Mg added to excess HCl produces 0.21 moles of $\text{H}_{2(g)}$ or 0.4g of $\text{H}_{2(g)}$

- Describe how you would make 100. mL of 2.0M NaNO_3 from powdered NaNO_3 . In the box, show what happens to the NaNO_3 when it is added to water.

- Determine mass of NaNO_3 needed:
 $2.0\text{M} = (\text{X g}/85.0 \text{ g/mol})/0.100 \text{ L}$ $\text{X} = 17\text{g}$
- Put 17g of NaNO_3 in a 100 mL volumetric flask
- Add water to 100 mL
- Mix



*Water molecules not shown for clarity of ions

Study Guide Redox Reaction

Martin Silberberg, Libby Weberg



Study Guide Redox Reaction:

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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02. Study Guide to Accompany Basics for Chemistry is an 18-chapter text designed to be used with Basics for Chemistry textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross-reference. The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill-in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter, and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and multiple-choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry, this book goes on exploring the concept of matter, energy, measurement, problem solving, atom, periodic table, and chemical bonding. These topics are followed by discussions on writing names and formulas of compounds, chemical formulas, and the mole, chemical reactions, calculations based on equations, gases, and the properties of a liquid. The remaining chapters examine the solutions, acids, bases, salts, oxidation-reduction reactions, electrochemistry, chemical kinetics, and equilibrium, and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry teachers and students.

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03. The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter, which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and

Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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