

Study Guide – Chapter 3: Stoichiometry

Test Date: Thursday, 12/18/14

- **Chemical Equations**
 - Writing word equations
 - remember to balance & include the states of matter
 - Balancing chemical equations w/coefficients
 - 5 types of reactions
 - Combination/synthesis, Decomposition, Combustion, Single Replacement, Double Replacement
 - identify what type of reaction is happening from a given equation
 - predict the products when given the reactants or told the type of reaction...
 - $C_3H_8 + O_2 (g)$ combustion forms products of $CO_2(g)$ and $H_2O (g)$
 - Synthesis of metal + nonmetal forms an ionic compound (criss cross charges for formula)
 - Decomposition breaks compound into its elements (watch out for diatomic BrINClHOFs or P_4 / S_8 solids)
- **Formula Weight (unit: amu) = Molar Mass (unit: g/mol)**
 - Calculations of formula weights and/or molar masses
 - % by mass = % composition calculations
- **Determining Empirical & Molecular Formulas**
 - Know the difference between an empirical and a molecular formula
 - Three types of problems
 - Normal problems will give % composition data or mass data to start
 - Combustion Analysis (like the quiz)
 - Hydrate Determination
 - Use the rhyme to help you remember the steps!
- **Stoichiometric Calculations**
 - Dimensional Analysis!!!
 - Avogadro's number (use for converting to or from # of atoms/molecules) → after molecules you can go to specific atom within the compound by looking at the chemical formula (ex: 4 H atoms per 1 molec. CH_4)
 - use the molar mass to go from mass to moles and/or moles to mass
 - Molar Ratios: use coefficients from a balanced chemical equation to go between any two different substances in a reaction (section 3.6)
- **Limiting Reactant Problems**
 - through calculations, identify which is the limiting reactant and which is the excess reactant
 - calculate the amount of product formed by the limiting reactant (this is the maximum or theoretical yield but you ALWAYS choose the LESSER quantity calculated)
 - determine how much of the excess reactant is used and how much is left over after the reaction is completed
 - calculate % yield
 - calculate theoretical yield and given the actual yield (from lab) of the substance
 - given % yield, you can determine actual yield collected in the lab

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