

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|---|---|--|--|
| <ul style="list-style-type: none"> • Condensation • Freezing • Vaporization • Melting • Boiling • Sublimation | <ul style="list-style-type: none"> • Deposition • Heat • Temperature • Thermochemistry • Exothermic • Endothermic | <ul style="list-style-type: none"> • Calorimeter • Joule • Specific heat • $q = m \cdot c \cdot \Delta T$ • $q = m \cdot \Delta H_{\text{vaporization}}$ • $q = m \cdot \Delta H_{\text{fusion}}$ | <ul style="list-style-type: none"> • Heat of reaction • Le Chatelier's principle • Equilibrium • Activation Energy • Catalyst |
|---|---|--|--|

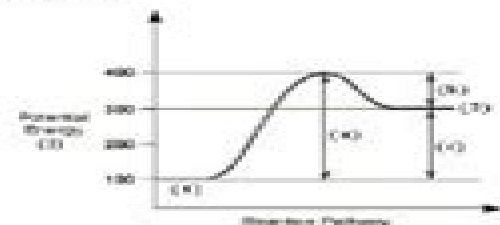
Book Assignments:

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$
 - Increase temp
 - Increase HCl
 - Increase pressure
 - Remove MgCl_2
 - Add H_2
- Identify the following terms with their correct number/letter from the graphs:
 - * Reactants- _____
 - * Products- _____
 - * Activation energy (number/letter & amount in Joules)- _____
 - * Is this reaction exothermic or endothermic?
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



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Resources in Education ,2000-10 **STOICHIOMETRY AND PROCESS CALCULATIONS** K. V. NARAYANAN,B. LAKSHMIKUTTY,2006-01-01 This textbook is designed for undergraduate courses in chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering safety engineering and industrial chemistry The chief objective of this text is to prepare students to make analysis of chemical processes through calculations and also to develop in them systematic problem solving skills The students are introduced not only to the application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and without chemical reactions The book presents the fundamentals of chemical engineering operations and processes in an accessible style to help the students gain a thorough understanding of chemical process calculations It also covers in detail the background materials such as units and conversions dimensional analysis and dimensionless groups property estimation P V T behaviour of fluids vapour pressure and phase equilibrium relationships humidity and saturation With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations Key Features SI units are used throughout the book Presents a thorough introduction to basic chemical engineering principles Provides many worked out examples and exercise problems with answers Objective type questions included at the end of the book serve as useful review material and also assist the students in preparing for competitive examinations such as GATE Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **Chemistry for the Health Professions** Charles H. Henrickson,Larry C. Byrd,1980 *Army Research and Development* ,1972 *Community and Junior College Journal* ,1976 Science Tests and Reviews Buros Center,1975 Science Tests and Reviews consisting of science sections of the first seven MMYs and Tests in Print II includes 217 original test reviews written by 81 specialists 18 excerpted test reviews 270 references on the construction use and validity of specific tests a bibliography on in print science tests references for specific tests cumulative name indexes for specific tests with references a publishers directory title index name index and a scanning index The 97 tests covered fall into the following categories 23 general 14 biology 35 chemistry 3 geology 6 miscellaneous and 16 physics

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