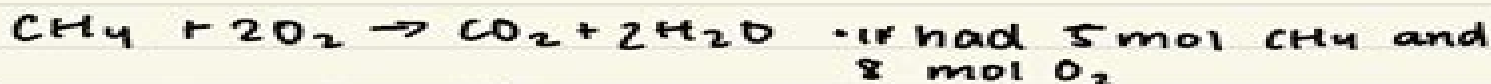


LIMITING REACTANT (thinked of grilled cheese)

- substance that is completely consumed in the chemical reaction, determines the amount of product that can be formed by reaction
 - ↳ identified by: # of moles of reactants, stoichiometry of balanced chemical equation
- when reactant runs out → reaction cannot proceed, dictates amount of product



$$5 \text{ mol CH}_4 \times \frac{2 \text{ mol O}_2}{1 \text{ mol CH}_4} = 10 \text{ mol O}_2$$

- you need 10 mol of O_2 to react with 5 mol of CH_4 but you only have 8 mol O_2 , so O_2 is the limiting reactant

$$8 \text{ mol O}_2 \times \frac{1 \text{ mol CH}_4}{2 \text{ mol O}_2} = 4 \text{ mol CH}_4 \quad \text{• you need 4 mol, have 5} \rightarrow \text{in excess}$$

LOOK AT MANY PRODUCTS

$$5 \text{ mol CH}_4 \times \frac{1 \text{ mol CO}_2}{1 \text{ mol CH}_4} = 5 \text{ mol CO}_2$$

$$8 \text{ mol O}_2 \times \frac{1 \text{ mol CO}_2}{2 \text{ mol O}_2} = 4 \text{ mol CO}_2$$

produces least amount of product so O_2 is the limiting reactant

also known as theoretical yield

THEORETICAL YIELD - calculated amount of product formed based on the amount of limiting reactant

ACTUAL YIELD - measured amount of product formed (what you get in the lab)

$$\text{PERCENT YIELD} = \frac{\text{ACTUAL YIELD}}{\text{THEORETICAL YIELD}} \times 100\%$$

Study Guide Stoichiometry Limiting Reactant Percent Yield

**Barron's Educational Series, Mark
Kernion, Joseph A. Mascetta**



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