

# Stoichiometry

## Answer Key

**Vocabulary:** Avogadro's number, balanced equation, cancel, coefficient, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

*[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]*

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? *176 grams of sugar*
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? *15 sticks of butter*

### Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo™ allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

$\text{Fe}_2\text{O}_3$  *1*

CO *3*

Fe *2*

$\text{CO}_2$  *3*

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

**Reactants**

Iron: *2*

Oxygen: *6*

Carbon: *3*

**Products**

Iron: *2*

Oxygen: *6*

Carbon: *3*

New Equation

$$\text{Fe}_2\text{O}_3 (\text{s}) + 3\text{CO} (\text{g}) \rightarrow 2\text{Fe} (\text{s}) + 3\text{CO}_2 (\text{g})$$

New Equation

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III)-oxide ( $\text{Fe}_2\text{O}_3$ )?

Problems solved: 0

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**Fredeen, Pearson**



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