

# PERIODIC TABLE (BASIC)

The periodic table is color-coded by groups: Noble gases (yellow), Halogens (orange), Chalcogens (green), Transition metals (blue), Alkali metals (pink), Alkaline earth metals (light blue), Lanthanides (light green), Actinides (light purple), and Other metals (dark blue). A temperature scale on the left ranges from 0 to 2000 K.

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	H	He															
2	Li	Be	B	C	N	O	F	Ne									
3	Na	Mg	Al	Si	P	S	Cl	Ar									
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I
6	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At
7	Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

For elements with no stable isotopes, the mass number of the isotope with the longest half-life is in parentheses.

## Hydrogen

Hydrogen is a chemical element with chemical symbol **H** and atomic number 1. With an atomic weight of 1.00794 u, hydrogen is the lightest element on the periodic table. Its electron configuration (1s<sup>1</sup>) is the most abundant chemical substance in the universe, constituting roughly 75% of all baryonic mass. Non-existence stars are mostly composed of hydrogen in its plasma state. The most common isotope of hydrogen, termed protium (named rarely used, symbol <sup>1</sup>H), has a single proton and one neutron.

## Helium

Helium is a chemical element with symbol **He** and atomic number 2. It is a colorless, odorless, tasteless, non-toxic, inert, monatomic gas that heads the noble gas group in the periodic table. Its boiling and melting points are the lowest among the elements and it exists only as a gas except in extreme conditions.

## Lithium

Lithium (from Greek: *lithos*, "stone", "stone") is a chemical element with symbol **Li** and atomic number 3. It is a soft, silver-white metal belonging to the alkali metal group of chemical elements. Under standard conditions it is the lightest metal and the least dense solid element. Like all alkali metals, lithium is highly reactive and flammable. For this reason, it is typically stored in mineral oil. When that option, lithium exhibits a metallic luster, but contact with moist air corrodes the surface quickly to a dull silvery gray, then black tarnish. Because of its high reactivity, lithium never occurs freely in nature, and instead, only appears in compounds, which are usually ionic. Lithium occurs in a number of igneous minerals, but due to its solubility as an ion is present in ocean water and is commonly obtained from brines and clays.

# Study Guides For Periodic Table

**Kathleen Armour**



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