

SECTION 2

INFORMATION BEFORE DRIVING YOUR TOYOTA

Information before driving your Toyota

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web jan 30 2020 ans the molality of solution is 1 mol/kg or 1 m example 02 11 11 g of urea NH_2CONH_2 was dissolved in 100 g of water calculate the molarity and molality of the solution given $n = 14$ $h = 1$ $c = 12$ $o = 16$ given mass of solute urea 11.11 g mass of solvent water 100 g 0.1 kg to find molarity of solution solution

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web solution moles of $\text{CH}_3\text{OH} = \frac{5.532}{32} = 0.173 \text{ mole}$ moles of $\text{H}_2\text{O} = \frac{40.18}{18} = 2.23 \text{ moles}$ therefore according to the equation mole fraction of $\text{CH}_3\text{OH} = \frac{0.173}{0.173 + 2.23} = 0.073$ to solve more examples on mole fraction formula and practice more question please visit byju s com

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is $0.210 + 0.050 = 0.260 \text{ atm}$ problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature

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