

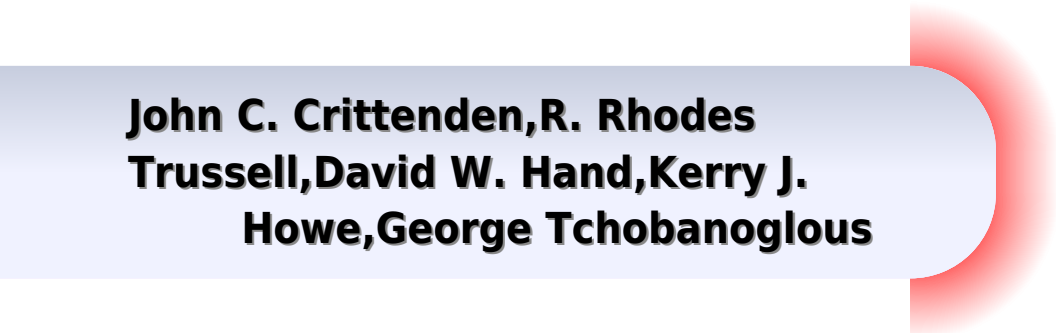
Topic

Water and Aqueous Systems

1. In the absence of any external forces, the shape of a drop of water in a vacuum is best described by which of the following?
- A. surface tension
 - B. density
 - C. viscosity
 - D. boiling point
2. The average molar mass of hydrous hematite (Fe_2O_3) is greater than that of anhydrous (Fe_2O_3) which is greater than the molar mass of the hydrous hematite-hydroxide ($\text{Fe}_2\text{O}_3 \cdot \text{H}_2\text{O}$) which is the sum of the molar mass of anhydrous hematite and water.
- A. water
 - B. hematite
 - C. hydrous hematite
3. Cobalt chloride (CoCl_2) is a blue-white solid and a hygroscopic solid and can be reduced to green cobaltous chloride. It dissolves in water to form a pink solution. Cobaltous chloride ($\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$) is the product of reduction of cobaltous chloride. Which of the following best describes cobaltous chloride?
- A. hydrous
 - B. anhydrous
 - C. cobaltous
 - D. cobaltic

Water And Aqueous Systems Practice Problems Answers

**John C. Crittenden,R. Rhodes
Trussell,David W. Hand,Kerry J.
Howe,George Tchobanoglous**



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A Problem-Solving Approach to Aquatic Chemistry James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter's main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses A Problem Solving Approach to Aquatic Chemistry serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings

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