

1. Balance the following equations. Then write **NET** ionic equations for each. Assuming 1L of total solution, calculate the concentration of each ion left in solution after reaction.



2. Calculate the number of:

- (a) grams in 0.25 mol of magnesium hydroxide. (14.5 g)
 (b) moles in 11.7 g of sodium chloride. (0.2 mol)
 (c) moles in 4.9 g of hydrogen sulphate. (0.05 mol)

3. Calculate the molarity of solutions containing the following amounts of solute:

- (a) 31.5 g HNO_3 per litre of solution. (0.5 M)
 (b) 26.5 g Na_2CO_3 per 200 mL of solution. (1.25 M)
 (c) 66.57 g AlCl_3 per 400 mL of solution. (1.25 M)

4. How many moles of solute are there in 100 mL of a 0.100 M solution? (0.01 mol)
 5. How many grams of NaCl are there in 250 mL of a 2.50 M NaCl solution? (36.6 g)
 6. What volume of a 0.75 M solution could be prepared from 500 g of Na_2SO_4 ? (4.7 L)
 7. One litre of 12 M HCl is diluted to 20 L. What is the molarity of the

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Charles Henrickson



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