Chapter 1

1.1

Mass of water $= 10^5$ g, temperature raised by $20°C$.
Heat needed $Q = 2 \times 10^7$ cal $= 8.37 \times 10^7$ J $= 23.2$ kwh.
Work needed $= mgh = 14 \times 150 \times 29000 = 6.09 \times 10^7$ ft-lb $= 22.9$ kwh.

1.2
Work done along various paths are as follows
ab:

$$\int_a^b P dV = N k_B T_1 \int_a^b \frac{dV}{V} = N k_B T_1 \ln \frac{V_b}{V_a}$$

cd:

$$P_2(V_d - V_b) = N k_B T_2 \left(1 - \frac{V_b}{V_d} \right)$$

de:

$$N k_B T_2 \int_d^e \frac{dV}{V} = N k_B T_2 \ln \frac{V_e}{V_d}$$

No work is done along bc and ea. The total work done is the sum of the above. Heat absorbed equals total work done, since internal energy is unchanged in a closed cycle.

1.3
(a)

$$\alpha = \frac{1}{V} \frac{\partial V}{\partial T} = \frac{b V_0 T^{b-1}}{T_0^b V}$$

(b)

$$\Delta V = \frac{b V_0 T^{b-1}}{T_0^b} \Delta T$$

$$P = \frac{N k_B T}{V} = \frac{N k_B T_0^b}{V_0} T^{1-b}$$

$$\text{Work done} = P \Delta V = b N k_B \Delta T$$

Solutions Manual Chler Introduction To Statistical Mechanics

Teunis C Dorlas

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Table of Contents Solutions Manual Chler Introduction To Statistical Mechanics

1. Understanding the eBook Solutions Manual Chler Introduction To Statistical Mechanics
    - The Rise of Digital Reading Solutions Manual Chler Introduction To Statistical Mechanics
    - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Manual Chler Introduction To Statistical Mechanics
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Solutions Manual Chler Introduction To Statistical Mechanics
    - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Manual Chler Introduction To Statistical Mechanics
    - Personalized Recommendations
    - Solutions Manual Chler Introduction To Statistical Mechanics User Reviews and Ratings
    - Solutions Manual Chler Introduction To Statistical Mechanics and Bestseller Lists
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