

Chemical Equilibrium

Ex 1. For the equilibrium: $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$

the equilibrium constant K_c is 50.5 at 448°C.

If $[\text{H}_2]_0 = [\text{I}_2]_0 = 1 \text{ mol.L}^{-1}$, what will the concentration of H_2 , I_2 and HI be at equilibrium.

Ex 2. Consider the following reversible reaction:



One mole of PCl_5 is placed in a flask of capacity 4 liter to be dissociated at constant temperature T. the number of moles of Cl_2 formed at equilibrium is 0.25mole.

Calculate the equilibrium constant of the above system at temperature T.

Ex 3. At 100°C, the equilibrium constant for the reaction system,



is $K_c = 2.19 \times 10^{-10}$

1. Determine whether the following mixture of components of the above system is at equilibrium:

$$[\text{COCl}_2] = 3.5 \times 10^{-3} \text{ mol.L}^{-1}$$

$$[\text{CO}] = 1.2 \times 10^{-5} \text{ mol.L}^{-1}$$

$$[\text{Cl}_2] = 3.3 \times 10^{-6} \text{ mol.L}^{-1}$$

2. If it is not at equilibrium, in what direction the system proceeds to reach equilibrium?

Ex 4. In a closed container of volume V and at constant temperature $T = 300\text{K}$, we introduce 2 mol of NO and 1 mol of O_2 gases. the following equilibrium is established:



Calculate the degree of conversion (α) of NO if the amount of NO_2 at equilibrium is three time that of O_2 .

Ex 5. Ethanamine $\text{C}_2\text{H}_5\text{NH}_2$ gas is put in an empty flask of volume $V = 2\text{L}$, maintained at constant temperature $T = 500\text{K}$.

The following equilibrium is established:



The initial amount of ethanamine is 0.2 mol and the total amount of the mixture at equilibrium is 0.3 mol.

1. Calculate the degree of dissociation α of ethanamine.
2. Calculate the equilibrium constant K_c at 500K.

Solutions Manual Chemical Equilibrium Chemistry

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**



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