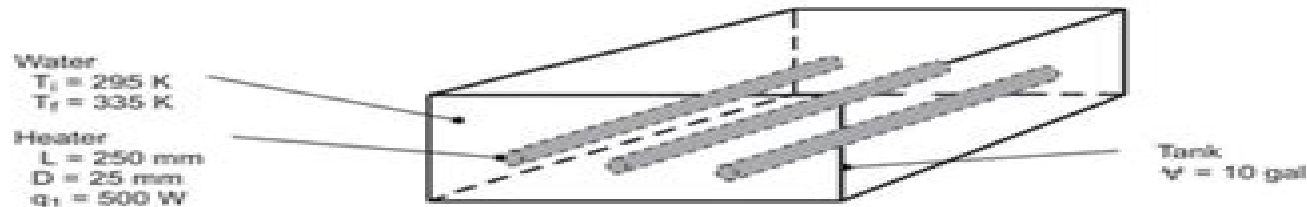


PROBLEM 1.38

KNOWN: Initial temperature of water and tank volume. Power dissipation, emissivity, length and diameter of submerged heaters. Expressions for convection coefficient associated with natural convection in water and air.

FIND: (a) Time to raise temperature of water to prescribed value, (b) Heater temperature shortly after activation and at conclusion of process, (c) Heater temperature if activated in air.

SCHEMATIC:



ASSUMPTIONS: (1) Negligible heat loss from tank to surroundings, (2) Water is *well-mixed* (at a uniform, but time varying temperature) during heating, (3) Negligible changes in thermal energy storage for heaters, (4) Constant properties, (5) Surroundings afforded by tank wall are large relative to heaters.

ANALYSIS: (a) Application of conservation of energy to a closed system (the water) at an instant, Eq. (1.11d), yields

$$\frac{dU}{dt} = Mc \frac{dT}{dt} = \rho V c \frac{dT}{dt} = \dot{q} = 3q_1$$

Hence,
$$\int_0^t dt = (\rho V c / 3q_1) \int_{T_i}^{T_f} dT$$

$$t = \frac{990 \text{ kg/m}^3 \times 10 \text{ gal} \left(3.79 \times 10^{-3} \text{ m}^3 / \text{gal} \right) 4180 \text{ J/kg} \cdot \text{K}}{3 \times 500 \text{ W}} (335 - 295) \text{ K} = 4180 \text{ s} \quad <$$

(b) From Eq. (1.3a), the heat rate by convection from each heater is

$$q_1 = Aq_1'' = Ah(T_s - T) = (\pi DL) 370(T_s - T)^{4/3}$$

Hence,

$$T_s = T + \left(\frac{q_1}{370\pi DL} \right)^{3/4} = T + \left(\frac{500 \text{ W}}{370 \text{ W/m}^2 \cdot \text{K}^{4/3} \times \pi \times 0.025 \text{ m} \times 0.250 \text{ m}} \right)^{3/4} = (T + 24) \text{ K}$$

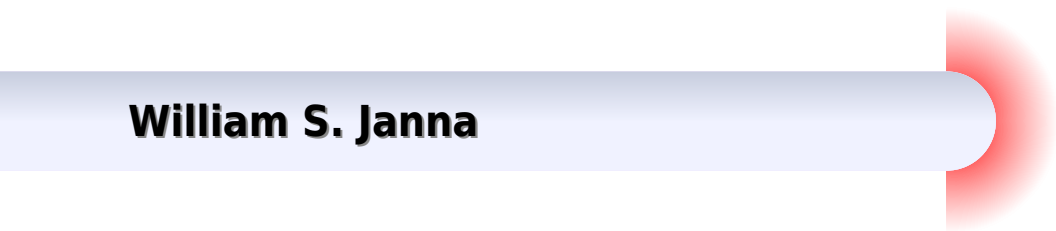
With water temperatures of $T_i = 295 \text{ K}$ and $T_f = 335 \text{ K}$ shortly after the start of heating and at the end of heating, respectively,

$$T_{s,i} = 319 \text{ K} \quad T_{s,f} = 359 \text{ K} \quad <$$

Continued

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