

Practice Problems: Acid-Base, Buffers

- In the titration of 80.0 mL of 0.150 M ethylamine, $\text{C}_2\text{H}_5\text{NH}_2$, with 0.100 M HCl, find the pH at each of the following points in the titration.
 - Initially, before any HCl has been added.
 - At the halfway point in the titration.
 - At the endpoint.
 - At 1/4 completion (the "one fourth of the way point")
 - At 3/4 completion
 - 10.0 mL of HCl beyond the endpoint.
- What volume of HCl is needed to reach each of the above points? (1 a-f above)
- What would be an appropriate indicator for the titration in #1?
- What volume of 0.120 M NaOH must be added to 100 mL of 0.100 M NaHC_2O_4 to reach a pH of 4.70?
- Describe the effect on the pH that results by adding
 - NaNO_2 to HNO_2
 - NaNO_3 to HNO_3 .
 Why are the effects not the same?
- A **buffer solution** is prepared by **dissolving 1.51 g of NH_3 and 3.85 g of $(\text{NH}_4)_2\text{SO}_4$ in 0.500 L of solution**.
 - What is the **pH of this solution**?
 - If **0.88 g of NaOH** is added, what is the pH?
- A buffer solution is prepared by dissolving **1.50 g each of benzoic acid, $\text{HC}_7\text{H}_5\text{O}_2$, and sodium benzoate, $\text{NaC}_7\text{H}_5\text{O}_2$, in 150.0 mL of solution**.
 - What is the **pH of this buffer solution**?
 - Which buffer component must be added, and in what quantity, to change the buffer pH to 4.00?
 - What quantity of 2.0 M NaOH or 2.0 M HCl must be added to the original buffer to change the pH to 4.00?
- In lab, you have the following chemicals available to you: 1.00 M NaOH, 1.00 M HCl, 1.00 M H_3PO_4 , NaH_2PO_4 (aq), Na_2HPO_4 (aq), and Na_3PO_4 (aq). Describe 3 different ways of making 200 mL of a buffer of pH = 12.10 in which the least concentrated buffer component is 0.10 M. Use only the above chemicals, plus distilled water.

ethylamine $\text{C}_2\text{H}_5\text{NH}_2$

$$K_b = 4.7 \times 10^{-4}$$

NH_3

$$K_b = 1.8 \times 10^{-5}$$

$\text{H}_2\text{C}_2\text{O}_4$

$$K_{a1} = 5.9 \times 10^{-2}$$

H_3PO_4

$$K_{a1} = 6.9 \times 10^{-3}$$

HC_2O_4^-

$$K_{a2} = 5.1 \times 10^{-5}$$

H_2PO_4^-

$$K_{a2} = 6.2 \times 10^{-8}$$

benzoic acid $\text{HC}_7\text{H}_5\text{O}_2$

$$K_a = 6.3 \times 10^{-5}$$

HPO_4^{2-}

$$K_{a3} = 4.8 \times 10^{-13}$$

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Siddappa N.Byrareddy



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