

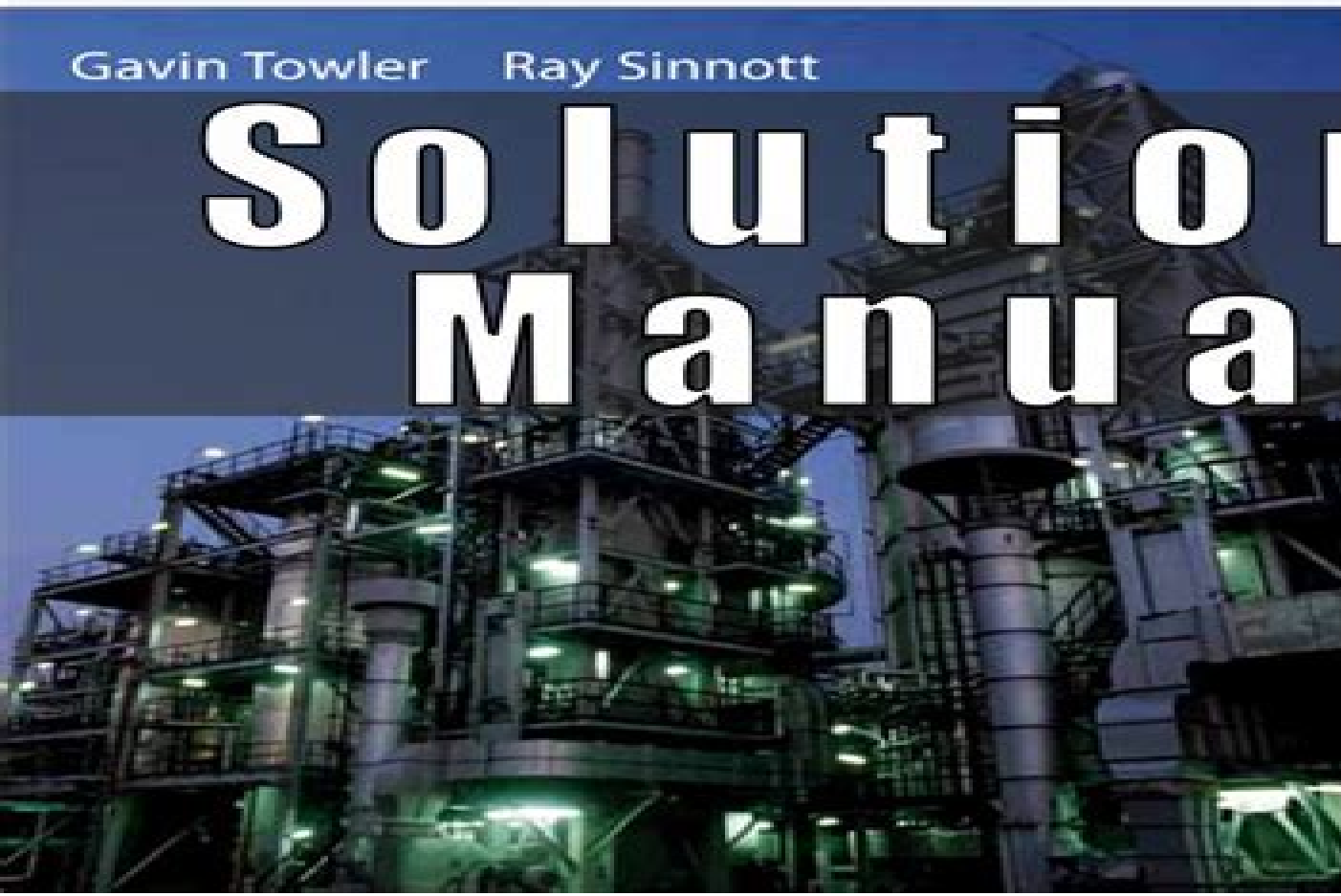
# CHEMICAL ENGINEERING DESIGN

Principles, Practice and Economics  
of Plant and Process Design

Second Edition

Gavin Towler Ray Sinnott

# Solutions Manual



# Towler 2nd Edition Solution Manual Chemical Design

**Xiang Xie**



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subsequent  $\Delta h$  values  $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{O}_2 \text{ (g)}$   $\Delta h = 2 \text{ H}_2 \text{ (g)} + \text{C}_2\text{H}_6 \text{ (g)}$   $\Delta h = 94.5 \text{ kJ}$   $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} \rightarrow \frac{1}{2} \text{O}_2 \text{ (g)}$   $\Delta h = 71.2 \text{ kJ}$   $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)} \rightarrow \text{CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (g)}$   $\Delta h = 283 \text{ kJ}$   $\text{C}_2\text{H}_6 \text{ (g)}$

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mole

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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