

Relative atomic mass of any element is measured using  $1/12^{\text{th}}$  mass of carbon isotope ( $C_{12}$ )

Relative- in relation with carbon.  
Usually fixed and mentioned

Molecular Mass- sum of the relative atomic masses of the elements present in a compound

## Stichometry:

A mole is that amount of chemical substance which contains as many particles (atoms, molecules or ions) as that present in 12 grams of carbon.

Avocadros Number: it is the number of particles that are present in one mole of a substance. It is numerically equal to  $6.022 \times 10^{23}$ . The symbol for it is

Relation between moles and number of particles:

Equation 1: Number of particles=number of moles \*  $N_A$  (avogadros constant)

Relation between moles and weight of the substance:

1 mole of any substance is equivalent to the molecular mass/molecular mass of the substance

Equation 2: Mass of a substance will be equal to the number of moles x the molecular mass

Relation between moles and volume (only for gasses)

1 mole of any gas is  $24\text{cm}^3$  at room temperature

$1\text{m}^3$  is 1000 meters

1 mol of oxy atoms=16 g of oxygen

# Target Stichiometry Lab

**Chao Zhang**



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