

**MULTIPLE CHOICE.** Choose the one alternative that best completes the statement or answers the question.

- 1) Consider the following reaction:



The average rate of appearance of B is given by  $\Delta[B]/\Delta t$ . Comparing the rate of appearance of B and the rate of disappearance of A, we get  $\Delta[B]/\Delta t = \underline{\hspace{1cm}} \times (-\Delta[A]/\Delta t)$ .

- A)  $-2/3$                       B)  $-3/2$                       C)  $+2/3$                       D)  $+3/2$                       E)  $+1$
- 2) A burning splint will burn more vigorously in pure oxygen than in air because
- A) nitrogen is a reactant in combustion and its low concentration in pure oxygen catalyzes the combustion.  
 B) oxygen is a reactant in combustion and the concentration of oxygen is higher in pure oxygen than it is in air.  
 C) oxygen is a product of combustion.  
 D) nitrogen is a product of combustion and the system reaches equilibrium at a lower temperature.  
 E) oxygen is a catalyst for combustion.
- 3) Of the following, all are valid units for a reaction rate except \_\_\_\_\_.
- A) mol/L                      B) M/s                      C) mol/hr                      D) mol/L-hr                      E) g/s

A flask is charged with 0.124 mol of A and allowed to react to form B according to the reaction  $A(g) \rightarrow B(g)$ . The following data are obtained for [A] as the reaction proceeds:

| Time (s)   | 1     | 10    | 20    | 30    | 40    |
|------------|-------|-------|-------|-------|-------|
| Moles of A | 0.124 | 0.110 | 0.088 | 0.073 | 0.054 |

- 4) The average rate of disappearance of A between 10 s and 20 s is \_\_\_\_\_ mol/s.
- A)  $1.1 \times 10^{-3}$                       B)  $4.4 \times 10^{-3}$                       C) 454                      D)  $9.90 \times 10^{-3}$                       E)  $2.2 \times 10^{-3}$

The peroxydisulfate ion ( $S_2O_8^{2-}$ ) reacts with the iodide ion in aqueous solution via the reaction:



An aqueous solution containing 0.050 M of  $S_2O_8^{2-}$  ion and 0.072 M of  $I^-$  is prepared, and the progress of the reaction followed by measuring  $[I^-]$ . The data obtained is given in the table below.

| Time (s)    | 0     | 400   | 800   | 1200  | 1600  |
|-------------|-------|-------|-------|-------|-------|
| $[I^-]$ (M) | 0.072 | 0.057 | 0.046 | 0.037 | 0.029 |

- 5) The concentration of  $S_2O_8^{2-}$  remaining at 800 s is \_\_\_\_\_ M.
- A) 0.015                      B) 0.041                      C) 0.076                      D)  $4.00 \times 10^{-3}$                       E) 0.046
- 6) Of the units below, \_\_\_\_\_ are appropriate for a first-order reaction rate constant.
- A)  $L \cdot mol^{-1} s^{-1}$                       B)  $s^{-1}$                       C)  $M^{-1} s^{-1}$                       D) mol/L                      E)  $M s^{-1}$

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