Multiple Choice: Acid/Base Titrations

1. A solution of known concentration is the definition of a

B) neutral solution. C) standard solution. D) saturated solution A) buffer solution.

AIMES CO.

 A weak acid (K_n = 4.5 x 10°) is neutralized with a weak base (K_n = 5.6 x 10°). What is the pH of the resulting solution?

A) acidic: B) basic G) neutrali D) pH = 3.2 E) pH = 7

ANS: A

 The equivalence point of a neutralization reaction is 8.5. What is the best indicator to use? A) crystal violet B) beomocresol green C) thymol blue D) methyl red E) alizatin

AMERICA.

 What is the pH of an aqueous solution that is 0.10 M HCOOH (K_s =1.8 x 10⁻⁶) and 0.10 M NaHCO₂? EN S. 7.4

B) 10.26 C) 5.62 D) 2.38 A) 3,74

ANS: A

1. Which of the following indicators would be most suitable for the titration of 0.10 M HBr(ag) with 0.10 M KOH(aq)?

A) bromothymol blue $(pK_{in} = 7.1)$ B) thymol blue $(pK_{in} = 1.7)$

C) bromophenol blue (pK_{in} = 3.9) D) alizarin vellow (p $K_h = 11.2$)

E) methyl orange (o.K., = 3.4).

ANS: A

 200.0 mL of 0.200 M HCl is titrated with 0.050 M NaOH. What is the pH after the addition of 100. mil. of the NaOH solution?

A) 0.93

Bb 1,45

CO 1.03

Do: 0.76

E) 0.82

ANS: A

1. If a 17.0 mL sample of 0.20 M sulfuric acid (H.SO.) is titrated with 0.14 M NaOH, it will require how many mL of titrant to reach neutralization?

A) 9.7 ml. B) 48.6 mL C) 389 mL D) 34.0 ml. E) 28.6 ml.

AINS: B

I. When NaOH(ao) is titrated with HCl(ag), what is the oH at the equivalence point?

A. equal to 7 B. greater than 7 C. less than 7 D. undefined

Titration Multiple Choice

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