

WEST BENGAL COUNCIL OF HIGHER SECONDARY EDUCATION
SYLLABUS FOR CLASSES XI AND XII
SUBJECT : CHEMISTRY (CHEM)

CLASS - XI

SEMESTER – I

SUBJECT : CHEMISTRY (CHEM)

FULL MARKS : 35

CONTACT HOURS : 70 Hours

COURSE CODE : THEORY

Sub-topics

UNIT No.	TOPICS	CONTACT HOURS	MARKS
Unit - 1	Some Basic Concepts of Chemistry: Laws of chemical combination. Concept of elements, atoms and molecules. Atomic and molecular masses. Mole concept and molar mass, percentage composition, empirical and molecular formula, chemical reactions, stoichiometry and calculations based on stoichiometry. Different concentration terms of solutions and related calculations.	07	03
Unit - 2	Structure of Atom: Bohr's model and its limitations, concept of shell and sub-shells, the dual nature of matter and light, de Broglie's relationship. Heisenberg uncertainty principle, Schrödinger wave equation (elementary idea only). Concept of orbitals, quantum numbers, shapes of s, p and d orbitals, rules for filling electrons in orbitals: Aufbau principle, Pauli exclusion principle and Hund's rule, exchange energy, electronic configuration of atom, stability of half-filled, completely filled orbitals.	12	06
Unit - 3	Classification of Elements and Periodicity in Properties: Modern periodic law and the present form of the periodic table, periodic trends in properties of elements – atomic radii, ionic radii, van der Waals' radii, ionization enthalpy, electron gain enthalpy, electronegativity, valency. Nomenclature of elements with atomic number greater than 100.	07	04
Unit - 4	Chemical Bonding and Molecular Structure: Valence electrons, ionic bond, bond parameters, covalent bond, Lewis structure, polar character of covalent bond, covalent character of ionic bond, valence bond theory, resonance, geometry of covalent molecules, VSEPR theory, concept of hybridisation, involving s, p and d orbitals and shapes of some simple molecules, intermolecular interactions, Hydrogen bonding. Molecular orbital theory of homonuclear diatomic molecules (H_2 , He_2 , O_2 , N_2 , F_2 – qualitative idea only)	13	06

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Table of Contents West Bengal Board Class Xii Science Syllabus

1. Understanding the eBook West Bengal Board Class Xii Science Syllabus
 - The Rise of Digital Reading West Bengal Board Class Xii Science Syllabus
 - Advantages of eBooks Over Traditional Books
2. Identifying West Bengal Board Class Xii Science Syllabus
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an West Bengal Board Class Xii Science Syllabus
 - User-Friendly Interface
4. Exploring eBook Recommendations from West Bengal Board Class Xii Science Syllabus
 - Personalized Recommendations
 - West Bengal Board Class Xii Science Syllabus User Reviews and Ratings
 - West Bengal Board Class Xii Science Syllabus and Bestseller Lists
5. Accessing West Bengal Board Class Xii Science Syllabus Free and Paid eBooks
 - West Bengal Board Class Xii Science Syllabus Public Domain eBooks
 - West Bengal Board Class Xii Science Syllabus eBook Subscription Services
 - West Bengal Board Class Xii Science Syllabus Budget-Friendly Options

6. Navigating West Bengal Board Class Xii Science Syllabus eBook Formats
 - ePub, PDF, MOBI, and More
 - West Bengal Board Class Xii Science Syllabus Compatibility with Devices
 - West Bengal Board Class Xii Science Syllabus Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of West Bengal Board Class Xii Science Syllabus
 - Highlighting and Note-Taking West Bengal Board Class Xii Science Syllabus
 - Interactive Elements West Bengal Board Class Xii Science Syllabus
8. Staying Engaged with West Bengal Board Class Xii Science Syllabus
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers West Bengal Board Class Xii Science Syllabus
9. Balancing eBooks and Physical Books West Bengal Board Class Xii Science Syllabus
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection West Bengal Board Class Xii Science Syllabus
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine West Bengal Board Class Xii Science Syllabus
 - Setting Reading Goals West Bengal Board Class Xii Science Syllabus
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of West Bengal Board Class Xii Science Syllabus
 - Fact-Checking eBook Content of West Bengal Board Class Xii Science Syllabus
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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