

Chemistry

Time: 2 hours

Marks: 60

1. Five isomeric para-disubstituted aromatic compounds A to E with molecular formula $C_6H_8O_2$ were given for identification. Based on the following observations, give structures of the compounds.
- (i) Both A and B form a silver mirror with Tollen's reagent; also B gives a positive test with $FeCl_3$ solution.
 - (ii) C gives positive iodoform test.
 - (iii) D is readily extracted in aqueous $NaHCO_3$ solution.
 - (iv) E on acid hydrolysis gives 1,4-dihydroxybenzene. [5]
2. 500 mL of 0.2 M aqueous solution of acetic acid is mixed with 500 mL of 0.2 M HCl at $25^\circ C$.
- (i) Calculate the degree of dissociation of acetic acid in the resulting solution and pH of the solution.
 - (ii) If 6 g of NaOH is added to the above solution, determine the final pH. [Assume there is no change in volume on mixing; K_a of acetic acid is $1.75 \times 10^{-5} \text{ mol L}^{-1}$]. [5]
3. Deduce the structures of $[NiCl_4]^{2-}$ and $[Ni(CN)_4]^{2-}$ considering the hybridization of the metal ion. Calculate the magnetic moment (spin only) of the species. [5]
4. The density of the vapour of a substance at 1 atm pressure and 500 K is 0.36 kg m^{-3} . The vapour effuses through a small hole at a rate of 1.33 times faster than oxygen under the same condition.
- (a) Determine (i) molecular weight, (ii) molar volume, (iii) compression factor (Z) of the vapour and (iv) which forces among the gas molecules are dominating, the attractive or the repulsive?
 - (b) If the vapour behaves ideally at 1000K, determine the average translational kinetic energy of a molecule. [5]
5. Write balance equations for the reactions of the following compounds with water;
- | | | |
|---------------|--------------|--------------|
| (i) Al_2C_3 | (ii) $CaNCN$ | (iii) BF_3 |
| (iv) NCl_3 | (v) XeF_4 | |
- [5]
6. Identify X, Y and Z in the following synthetic scheme and write their structures.
- $$CH_3CH_2C \equiv C-H \xrightarrow[\text{(ii) } CH_3CH_2Br]{\text{(i) } NaNH_2} X \xrightarrow{H_2/Pd-BaSO_4} Y \xrightarrow{\text{alkaline } KMnO_4} Z$$
- Is the compound Z optically active? Justify your answer. [5]
7. How is boron obtained from borax? Give chemical equations with reaction conditions. Write the structure of B_2H_6 and its reaction with HCl.

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Table of Contents What Is Chemistry Main Paper Exam Question For 2waec

1. Understanding the eBook What Is Chemistry Main Paper Exam Question For 2waec
 - The Rise of Digital Reading What Is Chemistry Main Paper Exam Question For 2waec
 - Advantages of eBooks Over Traditional Books
2. Identifying What Is Chemistry Main Paper Exam Question For 2waec
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an What Is Chemistry Main Paper Exam Question For 2waec
 - User-Friendly Interface
4. Exploring eBook Recommendations from What Is Chemistry Main Paper Exam Question For 2waec
 - Personalized Recommendations
 - What Is Chemistry Main Paper Exam Question For 2waec User Reviews and Ratings

- What Is Chemistry Main Paper Exam Question For 2waec and Bestseller Lists
- 5. Accessing What Is Chemistry Main Paper Exam Question For 2waec Free and Paid eBooks
 - What Is Chemistry Main Paper Exam Question For 2waec Public Domain eBooks
 - What Is Chemistry Main Paper Exam Question For 2waec eBook Subscription Services
 - What Is Chemistry Main Paper Exam Question For 2waec Budget-Friendly Options
- 6. Navigating What Is Chemistry Main Paper Exam Question For 2waec eBook Formats
 - ePub, PDF, MOBI, and More
 - What Is Chemistry Main Paper Exam Question For 2waec Compatibility with Devices
 - What Is Chemistry Main Paper Exam Question For 2waec Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of What Is Chemistry Main Paper Exam Question For 2waec
 - Highlighting and Note-Taking What Is Chemistry Main Paper Exam Question For 2waec
 - Interactive Elements What Is Chemistry Main Paper Exam Question For 2waec
- 8. Staying Engaged with What Is Chemistry Main Paper Exam Question For 2waec
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers What Is Chemistry Main Paper Exam Question For 2waec
- 9. Balancing eBooks and Physical Books What Is Chemistry Main Paper Exam Question For 2waec
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection What Is Chemistry Main Paper Exam Question For 2waec
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine What Is Chemistry Main Paper Exam Question For 2waec
 - Setting Reading Goals What Is Chemistry Main Paper Exam Question For 2waec
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of What Is Chemistry Main Paper Exam Question For 2waec
 - Fact-Checking eBook Content of What Is Chemistry Main Paper Exam Question For 2waec
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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