

Directions: The following is an End-Of-Course Review Guide designed to assist you as prepare for your EOC. It is imperative that you complete this guide to the best of your ability. This will help you to achieve a higher average on your second quarter grade.

Unit 1 & 2 Lab Safety, Introduction to Chemistry, & Measurement (Chapters 1 & 3):

1. What is chemistry?
2. Distinguish between hypothesis, theory, and law.
3. Classify the following as having good or poor accuracy and good or poor precision:
4. A scientist experimentally determines the speed of light to be 2.98×10^8 m/sec. In a second experiment, she determines the speed to be 2.99×10^8 m/sec.
5. The actual concentration of a solution is found to be 1.5 M. A scientist finds the concentrations to be 5.0 M and 5.2 M in two separate experiments.
6. Write the following in scientific notation:
a. 0.000 85 b. 1, 000, 000 c. 0.4565 d. 112
7. How many significant figures are in the following:
a. 0.98 b. 1, 000 c. 0.3450 d. 1.98×10^2
8. Report you answer to the correct number of significant figures:
a. $0.56 + 2.345$ b. $9.45 - 8.4$ c. 98.7×18.34 d. $1.23 / 3$
9. Explain how the measurement 45 cm is different from 45.0 cm.
10. List the SI units for the following quantities:
a. Mass b. Temperature c. Volume d. Length
11. How many meters are in the following:
a. 1 km b. 1 cm c. 1 mm
12. Calculate the density of a 5.0 g object that has a volume of 2.0 cm^3 .
13. Convert the following temperatures: a. 34°C to K b. 50 K to $^\circ\text{C}$
14. How will the graphs of a direct and indirect (inverse) proportion differ from each other?
15. Distinguish between periods and families.

Study Guide For Chemistry Eoc Tn

Padhraic Smyth



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Monthly Catalogue, United States Public Documents ,1984 *Scientific and Technical Aerospace Reports* ,1995
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Film and Video Finder ,1997 ,1997 Black Issues in Higher Education ,1984 *Reverse Acronyms, Initialisms, & Abbreviations Dictionary* ,1987 **Small Business Sourcebook** ,2010 **Reverse Acronyms, Initialisms & Abbreviations Dictionary**. Mary Rose Bonk,Regie Carlton,Gale Research Inc,1997 *Who's Who of American Women, 1986-1987* Marquis Who's Who,1986-11 **Cracking the Virginia SOL** Paul Foglino,Princeton Review (Firm),2000 The Princeton Review realizes that acing the Chemistry exam is very different from getting straight As in school The Princeton Review doesn't try to teach students everything there is to know about chemistry only the techniques they'll need to score higher on the exam There's a big difference In **Cracking the Virginia SOL** EOC Chemistry TPR will teach test takers how to think like the test makers and Score higher by mastering what test takers need to know with proven techniques Use word association to solve problems when unsure of the correct answers Ace the exam by understanding the logic of the test makers and applying it to solutions This book includes 2 full length simulated exams All of TPR's sample test questions are just like the ones test takers will see on the actual End of Course Chemistry exam and TPR fully explains every solution Contents Include The Mystery Exams About the SOL Exams How to Crack the Chemistry EOC II Chemistry Review Atomic Structure Anatomy of the Periodic Table Chemical Compounds Chemical Reactions Chemistry Math Gases Phase Changes Solutions Equilibrium Acids and Bases How Chemical Reactions Happen Laboratory III The Princeton Review Practice Tests
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