Covalent Bonding

How does a covalent bond form?

- A covalent bond forms when two atoms share one or more pairs of electrons.
- Atoms in covalent bonds typically have similar electronegativities and are usually nonmetals.
- The shared electrons allow each atom to achieve a full outer electron shell, similar to the noble gas configuration.

Types of Covalent Bonds

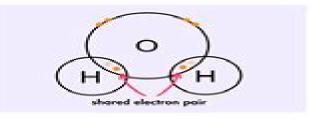
- · Single Bond: One pair of electrons is shared (e.g., H2, Cl2).
- Double Bond: Two pairs of electrons are shared (e.g., O₂, CO₂).
- Triple Bond: Three pairs of electrons are shared (e.g., N₂).



Dot-and-Cross Diagrams

- Only the outer electron shells (valence electrons) are shown.
- Shared pairs of electrons between atoms are shown as pairs of dots and crosses.
- It represents how the atoms bond and achieve stable electron configurations.

Example: Water H₂O



Properties of Covalent Compounds

- Covalent bonds CH,
- Melting and Boiling Points: Generally lower than ionic compounds because the forces between molecules are weaker.

discrete molecules with weak intermolecular

· Molecular Structure: They are made up of

- Electrical Conductivity: Poor conductors of electricity in all states, as they do not have free ions or electrons.
- Solubility: Often soluble in nonpolar solvents and may be insoluble in water, depending on their polarity.

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James E. Brady, Frederick A. Senese

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