

Chapter 4 The Energy of life

1. State two laws of Thermodynamics.
 - a. **First:** Energy is neither created nor destroyed. The total amount of energy is constant.
 - b. **Second:** Useful energy decreases when energy is converted from one form to another, and heat is produced
2. Define the following terms – Energy, Entropy, Kinetic energy, and potential energy.
 - a. **Energy:** the ability to do work
 - b. **Entropy:** increasing disorder due to useful energy becoming unusable heat
 - c. **Kinetic Energy:** the energy of position
 - d. **Potential Energy:** the energy of motion
3. Explain endergonic and exergonic reactions with examples.
 - a. **Endergonic:** reactions require the input of energy. Potential energy is less in the reactants than in the products. Ex: photosynthesis, building organic compounds
 - b. **Exergonic:** reactions release energy. Potential energy is more in the reactants than in the products. Ex: cellular respiration, breaking-down organic molecules
4. What is activation energy?
 - a. The minimum amount of energy needed to start a reaction
5. What are enzymes? List the factors that contribute to the enzyme substrate specificity?
 - a. Enzymes are biological catalysts that speed up reactions without being consumed. Substrates bind to an enzyme's active site.
 - b. Most enzymes are proteins
 - c. **Factors:** complementary shape, charge, and hydrophobic/hydrophilic character

Study Guide For Biology 160 Chapter 4

R Sanford



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