

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-Radical Reactions

There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Resonance Structures



Resonance Structures



There are two types of resonance structures

1. **Conjugation of double bonds**



2. **Conjugation of lone pairs**



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two multiple bonds consist of one sigma and two pi bonds.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **Polar bonds** are formed by the difference in electronegativity values of the two atoms.

• **Nonpolar bonds** are formed by the same electronegativity values of the two atoms.

• If the difference in electronegativity is between 0.5 and 1.7, the bond is polar. Examples are the C-Cl and C-Br bonds.

• If the difference in electronegativity is between 1.7 and 2.0, the bond is nonpolar. Examples are the C-C and C-H bonds.

• A difference of 2.0 or higher indicates that the electrons are not shared equally and a ionic bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The dipole moment is a vector pointing from the positive end to the negative end.

• The dipole moment is a vector pointing from the positive end to the negative end.

• The vector dipole moment is a vector pointing from the positive end to the negative end.

• The vector dipole moment is a vector pointing from the positive end to the negative end.

• μ (dipole moment) indicates an electron shift from one atom to another.

• μ (dipole moment) indicates an electron shift from one atom to another.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

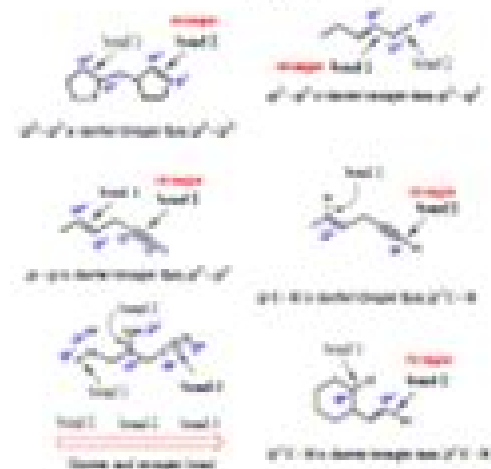


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	120 pm
Strength	356 kJ/mol	475 kJ/mol	591 kJ/mol

Higher s character means stronger bond

Examples



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