



$$\text{Moles (mol)} \times \text{Molar Mass (g/mol)} = \text{Mass (g)}$$

$$\text{Moles (mol)} \times \text{Avogadro's Number (atoms/mol)} = \text{atoms}$$

$$\frac{\text{Mass (g)}}{\text{Molar Mass (g/mol)}} = \text{Moles (mol)}$$

$$\frac{\text{Atoms}}{\text{Avogadro's Number (atoms/mol)}} = \text{Moles (mol)}$$

Study Guide Mixed Mole Conversions

Susan McMurry, John McMurry



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Study Guide for Chang's Chemistry Kenneth W. Watkins, 1988

Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney William H. Myers, William T. Mooney, 1980

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PPI FE Chemical Review Manual eText - 1 Year Michael R. Lindeburg, 2016-05-05

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