

Name _____ Date _____ Class _____

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

Solutions

Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	solution	solute	solution
insoluble	miscible	solute	solvent	

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen is air is the (2) solute and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air is (4) soluble substances. A substance that does not dissolve is (5) insoluble. (6) liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

Solution is the process of surrounding solute particles with solvent particles to form a solution.

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CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

True _____ 14. The overall energy change that occurs when a solution forms is called the heat of solution.

False _____ 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False _____ 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

True _____ 17. In a saturated solution, solvation and crystallization are in equilibrium.

True _____ 18. Additional solute can be dissolved in an unsaturated solution.

True _____ 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Business Communication Process and Product, Brief Edition, 7th Edition Mary Ellen Guffey, Dana Loewy, Esther Griffin, 2021-03-23 Business Communication Process and Product Seventh Brief Canadian Edition prepares students for a career in an increasingly digital and complex mobile social and global workplace With new Canadian case studies concept checks examples and references this new edition s content is even more relevant Students are introduced to the various recursive steps process in creating effective business documents product Business Communication Process and Product was developed to equip students with skills that will meet their future employers expectations such as written and oral communication skills critical thinking and analytical reasoning and ethical decision making Guffey Loewy Griffin Your authoritative and trusted brand in Business Communication Cbdce certified billing & coding specialist exam prep 2025-2026 Elliot Spencer, 2025-07-28 This comprehensive and expertly crafted study guide CBDCE Certified Billing Coding Specialist Exam Prep 2025 2026 is the answer you ve been searching for Designed specifically for candidates who want to conquer the CBDCE exam with confidence this book combines in depth content review realistic practice questions and proven test taking strategies to transform your preparation experience With over 600 meticulously developed practice questions and detailed answer explanations you ll gain not only knowledge but the confidence to tackle every question with precision Are you overwhelmed by the sheer volume of material you need to master to pass the CBDCE Certified Billing Coding Specialist Exam Do you find yourself struggling to organize your study time effectively unsure which resources truly prepare you for success You re not alone The journey to becoming a certified billing and coding specialist is challenging filled with complex medical terminology ever changing coding guidelines and high stakes exam pressure Many aspiring professionals face the frustration of sifting through outdated or incomplete study materials leaving them underprepared and anxious on exam day This comprehensive and expertly crafted study guide CBDCE Certified Billing Coding Specialist Exam Prep 2025 2026 is the answer you ve been searching for Designed specifically for candidates who want to conquer the CBDCE exam with confidence this book combines in depth content review realistic practice questions and proven test taking strategies to transform your preparation experience With over 600 meticulously developed practice questions and detailed

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