

Chemistry 1 Term 1 exam sample problems by topic

physical or chemical change

- boiling water **P**
- burning wood **P**
- melting aluminum **P**
- separating sand from water **P**

element names/symbols - know element symbols from handout during 1st week of class

metric prefixes (centi-, milli-, kilo-)

- Convert 54 mm to m.
- Convert 9.2×10^{-5} cL to L.
- Convert 5.12 kg to g.

$$\frac{54 \text{ mm}}{1000 \text{ mm}} = 0.054 \text{ m}$$

$$\frac{9.2 \times 10^{-5} \text{ cL}}{100 \text{ cL}} = 9.2 \times 10^{-7} \text{ L}$$

$$\frac{5.12 \text{ kg}}{1000 \text{ g}} = 5120 \text{ g}$$

Temperature conversions

- Convert 23 degrees C to Kelvin
- Convert 50 K to degrees C
- Convert 37 degrees C to degrees F

$$23 + 273 = 296 \text{ K}$$

$$50 - 273 = -223 \text{ C}$$

$$\frac{9}{5}(37 \text{ C}) + 32 = 98.6 \text{ F}$$

significant digits

- Round 8.9456 to 3 significant digits.
- Round 513431 to 2 significant digits.
- Add $4.324 + 345.1$, and round your answer properly.
- Multiply 32.56×103 , and round your answer properly.

$$8.95$$

$$510000$$

$$349.424 \rightarrow 349.4 \text{ (1 decimal place)}$$

$$3353.68 \rightarrow 3350 \text{ (3 sig figs)}$$

scientific notation

- Express 3712 in scientific notation.
- Express 0.0377561 in scientific notation.

$$3.712 \times 10^3$$

$$3.77561 \times 10^{-2}$$

density

What is the volume of a solid with a density of 7.8 g/mL and a mass of 11.8 g?

$$D = \frac{m}{V} \quad 7.8 \text{ g/mL} = \frac{11.8 \text{ g}}{V}$$

$$V = 1.51 \text{ mL}$$

unit conversions (inches to cm, seconds to hours, etc) including imperial and metric

- Convert 54 hours to seconds
- Convert 84 oz to grams
- Convert 1234 m to inches
- Convert 2.3 quarts to Liters

$$\frac{54 \text{ hr}}{1 \text{ hr}} \times \frac{60 \text{ min}}{1 \text{ min}} \times \frac{60 \text{ sec}}{1 \text{ min}} = 194400 \text{ sec}$$

$$\frac{84 \text{ oz}}{16 \text{ oz}} \times \frac{1 \text{ lb}}{1 \text{ lb}} \times \frac{454 \text{ g}}{1 \text{ lb}} = 2400 \text{ g}$$

$$\frac{2.3 \text{ qt}}{1.057 \text{ qt}} = 2.18 \text{ L}$$

$$\frac{1234 \text{ m}}{1 \text{ m}} \times \frac{100 \text{ cm}}{1 \text{ m}} \times \frac{1 \text{ inch}}{2.54 \text{ cm}} = 48600 \text{ inches}$$

Counting protons, neutrons, and electrons

Symbol	Protons	Neutrons	Electrons	Charge
${}^4_2\text{Be}$	4	5	4	0
${}^{47}_{19}\text{Ag}$	47	61	47	0
${}^{201}_{80}\text{Hg}$	80	121	80	0
${}^{33}_{16}\text{S}^{2-}$	16	17	18	-2
${}^{13}_{13}\text{Al}^{3+}$	13	15	10	+3
${}^{34}_{34}\text{Se}^{-}$	34	47	35	-1
${}^{60}_{28}\text{Ni}$	28	32	28	0
${}^{80}_{34}\text{Se}^{2-}$	34	46	36	-2
${}^{16}_8\text{O}^{2-}$	8	8	10	-2

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