

1. The Structure of the Atom

- The Particle Model states that all matter is made up of atoms.
- Atoms have origin in stars.
- We have been able to modify theories and develop new ones.

Nanomaterials

- Have diameters of 1-100 nm - nanometre as a useful unit.
- May have different properties than bulk materials.
- Large SA compared to V - surfaces are used to absorb and transport molecules as well as being catalysts for chemical reactions
- Measurements should be converted as below:

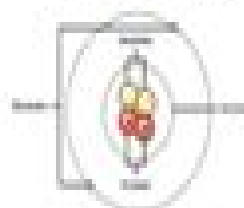
$$1 \text{ m} = 10^3 \text{ cm} = 10^3 \text{ mm} = 10^6 \mu\text{m} = 10^9 \text{ nm}$$

Historical development

1808	Dalton - elements have very small particles - atoms
1897	Thomson - positively charged spheres with negatively charged particles - electrons
1911	Rutherford - protons in the centre, electrons surrounding, mass concentrated in the nucleus
1913	Bohr - electrons orbit the nucleus in different energy levels, changing orbits - emit/gain energy
1926	Schrodinger - major electrons energy levels - subshells, orbitals are grouped together in the same subshell
1932	Chadwick - discovered the nucleus, explained the existence of isotopes

Location of particles

- Positively charged protons and uncharged neutrons have the same mass, both contribute the most to the atomic mass
- Negatively charged electrons move around the nucleus in specific energy levels, mass is apx. 1800 times smaller



Representing elements

- Atomic number (symbol Z) is the number of protons in the atom
- Mass number (symbol A) is the number of protons **and** neutrons
- In a neutral atom protons = electrons
- Mass number - Atomic number = number of neutrons
- Representing elements and isotopes is as below:



Isotopes

- Isotopes have same atomic number but different mass number
- isotopes have same chemical properties but different physical properties (mass, radioactivity, density)



Vce Chemistry Aspirin Sac

Robert P. Pinnell



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Aspirin, 2003 **How Aspirin Entered Our Medicine Cabinet** Steven M. Rooney, J.N. Campbell, 2017-02-17 This brief traces the story of one of our most common medicines aspirin On a journey involving science diverse characters shady business deals innovative advertising and good old fashioned luck Rooney and Campbell describe how aspirin was developed and marketed on a global scale Starting at the beginning of the twentieth century the authors explain the use of aspirin during the First World War the development of competition drugs such as ibuprofen during the interwar years and the application of aspirin to heart disease in the 1950s and 1960s On a broader level Rooney and Campbell show that the development of America's modern pharmaceuticals was a complex weaving of chemistry and mass culture They argue that aspirin's story provides a way to understand the application of complex chemical formulas in medical results This brief is of interest to historians of chemistry and medicine as well as the general educated reader Aspirin David Lewis, Colin Osborne, Maria J. Pack, 2003 Aimed at post 16 students in chemistry this book contains activities for both academic and vocational courses *Colorimetric Determination of Aspirin in Commercial Preparation* Robert P. Pinnell, 1989-01-01

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