SS 2 CHEMISTRY PRACTICAL 3RD TERM EXMINATION

 B is a solution containing 5.30g of anhydrous sodium trioxocarbonate (IV) per dm³. A is a solution of hydrochloric acid of unknown concentration.

Put A into the burette and titrate 25cm* (or 20cm*) portions of 8 with A. Using methyl orange as the indicator. Record the volume of your pipette.

Tabulate your burette readings, and calculate the average volume of acid used.

From your results, calculate:

- a. The concentration of solution B in moles per dm1
- b. The concentration of solution A in moles per dm3
- The mass concentration in g per dm³ of the acid in solution A.

The equation of reaction is:

$$2HCI+Na_2CO_3 \rightarrow 2NaCI+CO_2+H_2O$$

- (a). If you were provided with anhydrous CaCO₃, spatula and stirrer (i) list three other materials
 you would require to prepare a standard solution of CaCO₃ (ii) State what you would observe on
 adding dilute HCl to a portion of the CaCO₃
 - (b). State what would be observed when aqueous ammonia is added to: (i) Litmus paper; (ii) Pb(NO₁)₂ solution in drops until in excess.

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Pandey O.P. & Bajpai D.N. & Giri S.

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