



MARIST SECONDARY SCHOOL

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2021 MSCE RESULTS

NO	SUR NAME	FIRST NAME	AGRIC.	BIO	BUS	CHEM	CHIC	COMPU.	ENGLISH	FRENCH	GEOGRA	HISTORY	MATHS	ADD. MA	PHYSICS	AGGR.
1	GOMBACHIKA	SIBUSISO MARK	2	2		2	7	2	3		2		1	1	2	10
2	VINCENT	VINTHENG	3	3	3	2	6		3		1		1	1	2	10
3	CHAYEKHA	MURHENDERE MIGUEL	2	2	2	2	5		1		2	2	2		3	11
4	MAPANJE	NOEL	2	3		2	4	1	2		2		2	2	4	11
5	MWALILINO	PROMISE	2	2		3	5	2	5		2		2	1	3	11
6	MWENEBANDA	BENTRY PETER	2	3	2	2	5		3		2		2	1	3	11
7	MZIKAMANDA	CHIMWEMWE	2	3	2	2	5		2		1	2	5		4	11
8	CHALE	MICHAEL DAVID	2	3	2	2	4		2		2		2	2	3	12
9	CHIKOPA	INNOCENT BRAVE JN	2	3	2	2	4		2		2	2	3		4	12
10	CHINYAMU	WISDOM	3	2		2	6	2	3		2		2		2	12
11	CHITEKWE	TAKONDWA JONATH	2	3		2	3	2	2	4	2		3	2	3	12
12	KAKOTA	MADALITSO MARTIN	2	2		2	4	2	3		2	2	4		3	12
13	KALONGA	AKUZIKE	2	2	3	2	7		2		2	2	2		2	12
14	KUYOKWA	MTENDE JOSEPH	2	3	2	2	5		4	6	2		2		2	12
15	MASOKA	HAROLD	2	2		2	6	3	4		2		2	2	2	12
16	MLENDU	MARVEL	2	3	2	2	4		3		2	2	2		3	12
17	MTONGA	VITUMBIKO	2	2		2	4	2	2		2		2	2	2	12
18	MZUMARA	ISAIAH	2	2		2	6	2	3		2	2	3		3	12
19	NYIRENDA	ZAITHWA	2	3		2	5	2	4		2	2	2		2	12
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21	TSIRIZENI	HORRICE	2	2		2	6	3	2	4	2		2		2	12
22	CHIMOMBO	MAFUNO	2	4		2	5	4	5		2	2	2		3	13
23	CHIUNDIRA	STEPHEN	2	3		3	7	2	3		2		2	2	3	13
24	HAUYA	GIFT	3	4	3	2	5		3		2	1	3	2	4	13
25	KAYINGA	NATHAN	4	3		2	5	3	4		2	2	2		2	13
26	KAZIKA	JOSEPH CHITAMBALA	3	3	2	2	4		3		2	2	2		4	13
27	MACHESO	STEVEN	2	3	5	2	6		2		2	2	3		4	13
28	MALALU	CHANCY	3	3	2	2	6		3		2	2	2		3	13
29	MALAMULA	KELVIN FELIX	2	3	2	2	5		3	3	2		2		4	13
30	MAULUKA	GEORGE	4	2	2	4	7		3		2	2	4		2	13
31	MAWAYA	VINCENT	3	3	3	2	5		3		2		2	2	2	13

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Table of Contents St John Bosco Secondary Msce 2013 Results

1. Understanding the eBook St John Bosco Secondary Msce 2013 Results
 - The Rise of Digital Reading St John Bosco Secondary Msce 2013 Results
 - Advantages of eBooks Over Traditional Books
2. Identifying St John Bosco Secondary Msce 2013 Results
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an St John Bosco Secondary Msce 2013 Results
 - User-Friendly Interface
4. Exploring eBook Recommendations from St John Bosco Secondary Msce 2013 Results
 - Personalized Recommendations
 - St John Bosco Secondary Msce 2013 Results User Reviews and Ratings
 - St John Bosco Secondary Msce 2013 Results and Bestseller Lists
5. Accessing St John Bosco Secondary Msce 2013 Results Free and Paid eBooks
 - St John Bosco Secondary Msce 2013 Results Public Domain eBooks
 - St John Bosco Secondary Msce 2013 Results eBook Subscription Services
 - St John Bosco Secondary Msce 2013 Results Budget-Friendly Options
6. Navigating St John Bosco Secondary Msce 2013 Results eBook Formats

- ePub, PDF, MOBI, and More
- St John Bosco Secondary Msce 2013 Results Compatibility with Devices
- St John Bosco Secondary Msce 2013 Results Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of St John Bosco Secondary Msce 2013 Results
 - Highlighting and Note-Taking St John Bosco Secondary Msce 2013 Results
 - Interactive Elements St John Bosco Secondary Msce 2013 Results
- 8. Staying Engaged with St John Bosco Secondary Msce 2013 Results
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers St John Bosco Secondary Msce 2013 Results
- 9. Balancing eBooks and Physical Books St John Bosco Secondary Msce 2013 Results
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection St John Bosco Secondary Msce 2013 Results
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine St John Bosco Secondary Msce 2013 Results
 - Setting Reading Goals St John Bosco Secondary Msce 2013 Results
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of St John Bosco Secondary Msce 2013 Results
 - Fact-Checking eBook Content of St John Bosco Secondary Msce 2013 Results
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with

solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be ksp Ca(OH)_2 from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determining of calcium hydroxide the purpose of this was to titrate a saturated Ca(OH)_2 solution with a standard hcl solution and find the ksp or the oh concentration was found and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochloric

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web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form lution of calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for calcium was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web calculate the ksp for calcium hydroxide Ca^{2+} OH^{-} 2.0×10^{-2} ksp 5×10^{-4} find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-6}

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web 1 thalli are present in the form of colony 2 ball like colony is enveloped by a gelatinous sheath figs 56 57 3 balls are greenish to bluish green in colour

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