

## 2-4: Precipitation Reactions

1. Start *Virtual ChemLab*, select *Reactions and Stoichiometry*, and then select *Precipitation Reactions* from the list of assignments. The lab will open in the Inorganic laboratory.
2. React each of the cations (across the top) with each of the anions (down the left) according to the data table using the following procedures:

**Data Table**

|                                                 | $\text{AgNO}_3$ ( $\text{Ag}^+$ ) | $\text{Pb}(\text{NO}_3)_2$ ( $\text{Pb}^{2+}$ ) | $\text{Ca}(\text{NO}_3)_2$ ( $\text{Ca}^{2+}$ ) |
|-------------------------------------------------|-----------------------------------|-------------------------------------------------|-------------------------------------------------|
| $\text{Na}_2\text{CO}_3$ ( $\text{CO}_3^{2-}$ ) | <i>a</i>                          | <i>f</i>                                        | <i>k</i>                                        |
| $\text{Na}_2\text{S}$ ( $\text{S}^{2-}$ )       | <i>b</i>                          | <i>g</i>                                        | <i>l</i>                                        |
| $\text{NaOH}$ ( $\text{OH}^-$ )                 | <i>c</i>                          | <i>h</i>                                        | <i>m</i>                                        |
| $\text{Na}_2\text{SO}_4$ ( $\text{SO}_4^{2-}$ ) | <i>d</i>                          | <i>i</i>                                        | <i>n</i>                                        |
| $\text{NaCl}$ ( $\text{Cl}^-$ )                 | <i>e</i>                          | <i>j</i>                                        | <i>o</i>                                        |

- a. Drag a test tube from the box and place it on the metal test tube stand. You can then click on the bottle of  $\text{Ag}^+$  ion solution on the shelf to add it to the test tube.
  - b. Click on the **Divide** button on the edge of the lab bench four times to make four additional test tubes containing  $\text{Ag}^+$ . With one test tube in the metal stand and four others in the blue rack, click on the  $\text{Na}_2\text{CO}_3$  bottle on the reagent shelf and observe what happens in the window at the top left. Record your observation in the table above. If the solution remains clear, record NR for no reaction. Drag this test tube to the red disposal bucket on the right.
  - c. Drag a second tube from the blue rack to the metal stand. Add  $\text{Na}_2\text{S}$ , record your observations and discard the tube. Continue with the third, fourth and fifth tube, but add  $\text{NaOH}$ ,  $\text{Na}_2\text{SO}_4$ , and  $\text{NaCl}$  respectively. Record your observations and discard the tubes. When you are finished, click on the red disposal bucket to clear the lab.
  - d. Repeat steps a-c for five test tubes of  $\text{Pb}^{2+}$  and  $\text{Ca}^{2+}$ . Record your observations in the data table. If no precipitate forms write NR for no reaction.
3. What happens in grid space d? What other reactions give similar results? Is it necessary to write an equation when no reaction occurs? Explain. \_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_
  4. Write balanced equations for all precipitation reactions you observed.
    - a.
    - b.
    - c.

# Virtual Lab Precipitation Reactions Answers

**Pedro Eduardo José Rivera Díaz del  
Castillo**



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