

# Oxidation-Reduction Worksheet 1

1. Identify the following in the reaction:  $2\text{Fe}^{+2} + \text{Cl}_2 \rightarrow 2\text{Fe}^{+3} + 2\text{Cl}^{-1}$  (+ marks)

a. The oxidizing agent	$\text{Cl}_2$	e. The species gaining electrons	$\text{Cl}_2$
b. The species being oxidized	$\text{Fe}^{2+}$	f. The species losing electrons	$\text{Fe}^{2+}$
c. The reducing agent	$\text{Fe}^{2+}$	g. The product of oxidation	$\text{Fe}^{3+}$
d. The species being reduced	$\text{Cl}_2$	h. The product of reduction	$\text{Cl}^{-}$

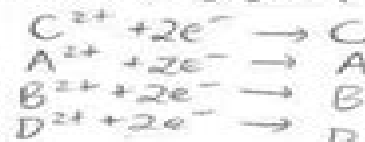
2. Use the Standard Reduction Table to answer the following. (+ marks)

- a. Will  $\text{Br}_2$  oxidize  $\text{Au}_{(s)}$ ? Explain. Yes,  $\text{Br}_{2(s)}$  is higher on the table as an oxidizing agent than  $\text{Au}_{(s)}$  as a reducing agent
- b. Give the symbol of an ion that will oxidize  $\text{Mn}_{(s)}$  but not  $\text{Cr}_{(s)}$ .  $\text{Zn}^{2+}$
- c. Which is a stronger reducing agent,  $\text{Sn}^{2+}$  or  $\text{Fe}^{2+}$ ?  $\text{Sn}^{2+}$
- d. Can you safely put a gold ring in acidified dichromate solution? What about acidified bromate solution? Explain.  $\text{Au}_{(s)}$  is higher on the table as a reducing agent than  $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+$  as an oxidizing agent, so it will not react.  $\text{Au}_{(s)}$  will not react with  $\text{BrO}_3^- + 6\text{H}^+$  either for same reason

3. Four metals A, B, C, & D were tested with separate solutions of  $\text{A}^{2+}$ ,  $\text{B}^{2+}$ ,  $\text{C}^{2+}$  &  $\text{D}^{2+}$ . Some of the results are summarized in the following table:

Metal	$\text{A}^{2+}$	$\text{B}^{2+}$	$\text{C}^{2+}$	$\text{D}^{2+}$
A		(I) no reaction	(II) reaction	
B				(III) no reaction
D	(IV) reaction			

List the ions in order from the strongest to weakest oxidizing agent. (2 marks)



4. Balance the following half reactions. (3 marks)

- a.  $\text{BrO}_3^- \rightarrow \text{Br}^-$  (basic)



- b.  $\text{HClO}_2 \rightarrow \text{ClO}_3^-$  (acidic)



- c.  $\text{Ga} \rightarrow \text{H}_2\text{GaO}_3^-$  (basic)



# Topic 9 Oxidation And Reduction Answer Key

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